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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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what is the limiting reagent if 8.52 g are actually formed what is the

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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ grams O}_2}{32 \text{ grams/mole}} = 0.6656 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
4 Al + 3 O₂ → 2 Al₂O₃
b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.926 \text{ mol Al}$
 $\frac{101.964 \text{ g Al}_2\text{O}_3}{302.3 \text{ g/mol}} = 0.337 \text{ mol Al}_2\text{O}_3$
c determine the percent yield
 $\frac{26.982 \text{ g Al}_2\text{O}_3}{101.964 \text{ g Al}_2\text{O}_3} \times 100 = 26.4\%$

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