

- (d) The equilibrium mixture leaving the reaction chamber contains 15% ammonia. Suggest how the ammonia could be separated from the mixture.

	boiling point/ $^{\circ}\text{C}$
hydrogen	-253
nitrogen	-196
ammonia	-33

[2]

- (e) Ammonia is used to make nitrogen trifluoride, NF_3 . Nitrogen trifluoride is essential to the electronics industry. It is made by the following reaction.



Determine if the above reaction is exothermic or endothermic using the following bond energies and by completing the following table. The first line has been done as an example. Bond energy is the amount of energy, in kJ/mole , needed to break or make one mole of the bond.

bond	bond energy in kJ/mole
$\text{N} - \text{H}$	390
$\text{F} - \text{F}$	155
$\text{N} - \text{F}$	280
$\text{H} - \text{F}$	565

bond	energy change/ kJ
$\text{N} - \text{H}$	$(3 \times 390) = 1170$
$\text{F} - \text{F}$	
$\text{N} - \text{F}$	
$\text{H} - \text{F}$	

[4]

[Total: 16]

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