

**Balancing a Redox Equation for an Acidic Solution  
Using the Oxidation Number Method**  
(Student textbook page 615)

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



**What Is Required?**

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

**What Is Given?**

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

| Plan Your Strategy   | Act on Your Strategy   |
|--|--|
| <b>Step 1</b> Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction.                                      | $\overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g})$ <p>The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.</p>   |
| <b>Step 2</b> Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number. | <p>Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the <math>\text{O}_2</math> lost 12 electrons and the <math>\text{CS}_2</math> gained 4 electrons.</p> $\begin{array}{c} \text{1}(-12\text{e}^-) = -12\text{e}^- \\ \downarrow \\ \overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g}) \\ \uparrow \qquad \qquad \qquad \uparrow \\ \text{3}(+4\text{e}^-) = +12\text{e}^- \end{array}$ |

# Solutions Manual Teacher Redox

**Andrew Burrows, John Holman, Simon  
Lancaster, Andrew Parsons, Tina  
Overton, Gwen Pilling, Gareth Price**



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