



Chapter 10 The Problem Solutions

1. The ion spacings and melting points of the sodium halides are as follows:

	NaF	NaCl	NaBr	NaI
Ion spacing, nm	0.23	0.28	0.29	0.32
Melting point, °C	988	801	740	660

Explain the regular variation in these quantities with halogen atomic number.

[Sol]

The halogenic atomic numbers are $Z = 9$ for fluorine (F), $Z = 17$ for chlorine (Cl), $Z = 35$ for bromine (Br) and $Z = 53$ for iodine (I). The greater the atomic number of a halogen atom, the larger the atom is, hence the increase in the interatomic spacing with Z . The larger the ion spacing, the smaller the cohesive energy, hence the lower the melting point.

3. (a) The ionization energy of potassium is 4.34 eV and the electron affinity of chlorine is 3.61 eV. The Madelung constant for the KCl structure is 1.748 and the distance between ions of opposite sign is 0.314 nm. On the basis of these data only, compute the cohesive energy of KCl. (b) The observed cohesive energy of KCl is 6.42 eV per ion pair. On the assumption that the difference between this figure and that obtained in (a) is due to the exclusion-principle repulsion, find the exponent n in the formula Br^{-n} for the potential energy arising from this source.

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