

Chapter 9 Review

- 1) What does stoichiometry mean? "measuring elements"
- 2) Before performing a stoichiometry problem, you must first do what with the chemical equation? **balance it**
- 3) The coefficients in a chemical equation represent what? **moles**
- 4) In a chemical equation, what relationships are shown? **molar ratios**
- 5) Solve the following mole-mole conversions using this equation:



- a) If 3.5 moles of sodium bicarbonate are decomposed, how many moles of carbon dioxide are produced?

$$3.5 \text{ mol NaHCO}_3 \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} = 1.8 \text{ mol CO}_2$$

- b) If 1.29 moles of water are produced by the decomposition of sodium bicarbonate, how many moles of sodium carbonate are produced?

$$1.29 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{1 \text{ mol H}_2\text{O}} = 1.29 \text{ mol Na}_2\text{CO}_3$$

- c) How many moles of sodium bicarbonate are required to produce 19.5 moles of carbon dioxide?

$$19.5 \text{ mol CO}_2 \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol CO}_2} = 39.0 \text{ mol NaHCO}_3$$

- 6) Use the equation in # 5 to solve the following mass-mass conversions.

- a) If given 10.5 grams of sodium bicarbonate, how many grams of sodium carbonate can be produced?

$$10.5 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{2 \text{ mol NaHCO}_3} \times \frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} = 6.62 \text{ g Na}_2\text{CO}_3$$

- b) How many grams of sodium bicarbonate would be required to produce 9.0 grams of water?

$$9.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol H}_2\text{O}} \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3} = 84 \text{ g NaHCO}_3$$

- 7) What conversion factor (mole ratio) would you use to solve for moles of sulfur formed from sulfur dioxide in the reaction



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Table of Contents Solving Problems A Chemistry H Answer Key

1. Understanding the eBook Solving Problems A Chemistry H Answer Key
 - The Rise of Digital Reading Solving Problems A Chemistry H Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Solving Problems A Chemistry H Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solving Problems A Chemistry H Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solving Problems A Chemistry H Answer Key

- Personalized Recommendations
 - Solving Problems A Chemistry H Answer Key User Reviews and Ratings
 - Solving Problems A Chemistry H Answer Key and Bestseller Lists
5. Accessing Solving Problems A Chemistry H Answer Key Free and Paid eBooks
- Solving Problems A Chemistry H Answer Key Public Domain eBooks
 - Solving Problems A Chemistry H Answer Key eBook Subscription Services
 - Solving Problems A Chemistry H Answer Key Budget-Friendly Options
6. Navigating Solving Problems A Chemistry H Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Solving Problems A Chemistry H Answer Key Compatibility with Devices
 - Solving Problems A Chemistry H Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Solving Problems A Chemistry H Answer Key
 - Highlighting and Note-Taking Solving Problems A Chemistry H Answer Key
 - Interactive Elements Solving Problems A Chemistry H Answer Key
8. Staying Engaged with Solving Problems A Chemistry H Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solving Problems A Chemistry H Answer Key
9. Balancing eBooks and Physical Books Solving Problems A Chemistry H Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solving Problems A Chemistry H Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solving Problems A Chemistry H Answer Key
- Setting Reading Goals Solving Problems A Chemistry H Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solving Problems A Chemistry H Answer Key

- Fact-Checking eBook Content of Solving Problems A Chemistry H Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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