



# Southbend 448ee 3gl Ranges Owners Manual

**Scott C. Dulebohn**



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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b = 1$  2 the multiplication of the mole fraction by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a$  mol a mol a mol b and  $x_b$  mol b mol a mol b note in the given mixture the sum of all the mole fractions present is equal to one  $x_a + x_b = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 30 g benzene s molar mass  $C_6H_6$  128 g mol<sup>-1</sup> hence  $C_6H_6$  s number of moles  $\frac{30}{128}$  mol 0 2344 mol molar mass of carbon tet  $CCl_4$  154 g mol<sup>-1</sup>

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$   $d = 0.660 \text{ g cm}^{-3}$  when it is mixed with heptane  $C_7H_{16}$   $d = 0.684 \text{ g cm}^{-3}$  in a 35.0 : 55.0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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of acetone in a solution if 19.0 ml of acetone was dissolved in 98.4 ml of xylene 7.1

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web since the mole fraction of the  $O_2$  is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature

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