

Answers to self-tests and exercises

Please note that just the final numerical solutions and short-answer solutions are supplied. For full worked solutions, please refer to the Solutions Manual by Alex Hadenic (97804706781712).

CHAPTER 1

Self-tests

- SL.1 759867cm^{-1} , $1.28 \times 10^{-7}\text{m}$ or 1280-nm .
 SL.2 3d set of orbitals, 5 orbitals.
 SL.3 d.
 SL.4 3p.
 SL.5 The s electron already present in Li repels the incoming electron more strongly than the p electron already present in B repels the incoming p electron, because the incoming p electron goes into a new orbital.
 SL.6 $\text{Ni} : [\text{Ar}][\text{d}^8 4s^2]$, $\text{Ni}^{2+} : [\text{Ar}][\text{d}^8]$
 SL.7 Period 4, Group 2, s block.
 SL.8 Generally, going down a group the atomic radius increases and the first ionization energy generally decreases.
 SL.9 Group 14.
 SL.10 The electron-gain process for N is unfavourable.
 SL.11 Na^+

Exercises

- E.1 $J(1s^2)(2s2p^2) = J^2(2s^2)(2p^2)(4s^2) = 2^2/4^2 = 0.25$
 E.2 (a) Ar nucleus
 (b) Exactly at Bohr radius, a_0 .
 (c) $3a_0 = 4.76\text{\AA}$.
 E.3 $J(2s, n=1) = J(2p, n=0) = 13.6\text{-eV} = 0.510\text{ a.u.}$
 $\text{eV} = 1.3 \times 10^{-19}\text{J}$.
 E.4 0.544eV . The discrepancy is due to the shielding effect.
 E.5 14.0eV , 4.16 eV

- E.6 102.5 nm , $1.937 \times 10^{-18}\text{ J}$
 E.7 $1.526 \times 10^{-15}\text{m}^{-1}$, 656.3 nm
 E.8 $\frac{1}{\lambda} = R\left(\frac{1}{1^2} - \frac{1}{\infty^2}\right) = 1.0974 \times 10^7\text{ m}^{-1}$
 $\frac{1}{\lambda} = R\left(\frac{1}{1^2} - \frac{1}{4^2}\right) = 1.0288 \times 10^7\text{ m}^{-1}$
 $\frac{1}{\lambda} = R\left(\frac{1}{1^2} - \frac{1}{9^2}\right) = 9.7547 \times 10^6\text{ m}^{-1}$
 $\frac{1}{\lambda} = R\left(\frac{1}{1^2} - \frac{1}{25^2}\right) = 8.2305 \times 10^6\text{ m}^{-1}$
 E.9 For a given value of n , the angular momentum quantum number l can assume all integer values from 0 to $n-1$.
 E.10 m^2 (a.g., $m^2 = 0$ for $n = 1$, $m^2 = 4$ for $n = 2$, etc.)

E.11

n	l	m_l	Orbital designation	Number of orbitals
1	0	+1, 0, -1	2p	3
2	1	+2, -1, 0, -2	3d	5
3	2	0	4s	1
3	2	+3, -2, -1, 0, -3	4f	7

- E.12 When $n = 5$, $l = 0$ (for the 1 orbitals) and $m_l = -3, -2, -1, 0, 1, 2, 3$, which represent the seven orbitals that complete the 5d subshell. The 5d orbitals represent the start of the actinoids, starting with Th and ending with La.
 E.13 Comparing the plots for 1s (Figure 1.10b) and 2s (Figure 1.12) orbitals, the radial distribution function for a 1s orbital has a single maximum, and that for a 2s orbital has two maxima and a minimum (at $r = 2a_0/2$ for hydrogenic 2s orbitals). The presence of the node at $r = 2a_0/2$ for $40(2s)$ requires the presence of the two maxima and the minimum in the 2s radial distribution functions. The absence of a radial node for $40(2p)$ requires that the 2p radial distribution functions has only a single maximum.

Solutions Manual Inorganic Chemistry Huheey

Ian Pickup



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