

Esempio: calcolare quante molecole sono contenute in un bicchier d'acqua avente la capacità di 0,135 litri.

L'acqua presenta una densità di 1 kg dm^{-3} ($1 \text{ dm}^3 = 1 \text{ l}$) e 0,135 l pesano pertanto 135g che corrispondono a

$$n_{H_2O} = \frac{W_{H_2O}}{PM_{H_2O}} = \frac{135g}{18g / mol} = 7,5 mol$$

il numero di molecole presenti in 7,5 moli sarà dunque pari a $7,5 \text{ mol} \times 6,022 \cdot 10^{23} \text{ mol}^{-1} = 4,52 \cdot 10^{24}$

Una conseguenza del principio di Avogadro è che un medesimo numero di moli di una qualsiasi sostanza gassosa devono occupare sempre il medesimo volume (a P e T costanti). Infatti se *"volumi uguali di gas diversi nelle stesse condizioni di T e P contengono lo stesso numero di particelle"*, allora deve anche essere vero che *"gas che contengano lo stesso numero di particelle devono occupare lo stesso volume"*. Se ne deduce pertanto che 1 mole di un qualsiasi gas, contenendo sempre lo stesso numero di particelle (il numero di Avogadro) deve occupare sempre il medesimo volume ed in particolare, a 0°C e alla pressione di 1 atm occupa un volume pari a **22,414 l**, detto **volume molare standard**.

1.9 Calcolo delle quantità che reagiscono

Come abbiamo già avuto modo di dire il concetto di mole è essenziale per semplificare i calcoli quantitativi o stechiometrici. La **stechiometria** (dal greco *στοιχειον* = elemento, sostanza fondamentale) è quella parte della chimica che si occupa degli aspetti quantitativi delle reazioni ed in particolare dei **rapporti numerici e ponderali** che intercorrono tra le specie chimiche che reagiscono.

Per poter procedere con i calcoli stechiometrici è necessario che una equazione sia bilanciata.

In un'equazione bilanciata sono già definiti i rapporti numerici tra specie chimiche.

Quando scriviamo l'equazione bilanciata



individuiamo ad esempio il rapporto numerico di reazione tra Idrogeno e Azoto che risulta essere pari a 3:1. Il rapporto tra Idrogeno ed ammoniaca è invece di 3:2 e così via.

Ma i coefficienti stechiometrici non rappresentano solo il numero di molecole, ma anche il numero di moli che reagiscono. Per rendercene conto, dopo aver ricordato che un'eguaglianza rimane tale se moltiplichiamo entrambi i suoi membri per uno stesso numero, moltiplichiamo entrambi i membri per il numero di Avogadro (N_A).

$$N_A \cdot (3\text{H}_2 + \text{N}_2) = N_A \cdot (2\text{NH}_3)$$

Applichiamo la proprietà distributiva:

$$3 N_A \cdot (\text{H}_2) + 1 N_A \cdot (\text{N}_2) = 2 N_A \cdot (\text{NH}_3)$$

ma per la definizione di mole: un numero di Avogadro di molecole di Azoto costituiscono una mole di Azoto, 3 numeri di Avogadro di molecole di Idrogeno costituiscono 3 moli di Idrogeno e 2 numeri di Avogadro di molecole di Ammoniaca costituiscono 2 moli di Ammoniaca.

Se ne deduce quindi che i rapporti numerici precedentemente individuati tra i coefficienti stechiometrici non sono solo rapporti molecolari, ma anche rapporti molari. In altre parole per ogni 3 moli di Idrogeno reagisce 1 mole di Azoto per dare 2 moli di Ammoniaca.

Sommario Sommario Teoria Atomica E Leggi Quantitative

Raffaele Leonardi



Sommario Sommario Teoria Atomica E Leggi Quantitative:

Manuale della sicurezza dei gas tossici. Guida completa per la preparazione all'esame di abilitazione Calabrese Maurizio, Di Giuseppe Dina, Rossini Renato, 2017-02-28 Uno strumento fondamentale per affrontare l'esame di abilitazione e per gli operatori pubblici e privati che si occupano di sicurezza ed igiene del lavoro relativamente ai gas tossici. Il tema gas tossici pur se datato nella sua impostazione normativa originaria deve essere letto facendo riferimento al progressivo recepimento di norme comunitarie che sono state emanate negli ultimi decenni. Le Direttive comunitarie costituiscono nel campo della sicurezza una vera e propria rivoluzione che pur nelle innovazioni contenute ha lasciato in piedi molti istituti previgenti tra cui per l'appunto quello relativo al R.D. 147/27 sui gas tossici. La prima parte del manuale sostanzialmente di carattere generale con l'illustrazione delle disposizioni del Regio Decreto le novità regolamentari europee sulle sostanze chimiche REACH e CLP, il sistema organizzativo di gestione della sicurezza previsto dal D.Lgs. n. 81/08 e le modalità d'accertamento degli organi di vigilanza. La seconda parte tratta l'imballaggio e stoccaggio, l'impiego dei gas, delle norme riguardanti il trasporto ADR, la prevenzione incendi, la sorveglianza sanitaria e delle misure di primo soccorso, dei dispositivi di protezione individuale, l'applicazione della direttiva ATEX e della direttiva Seveso. Inoltre contiene utilissime schede suddivise per singolo gas tossico organizzate in maniera tale da essere strumento di consultazione e di lavoro immediato per tutti coloro che si occupano di gas tossici. Il CD Rom allegato contiene modulistica normativa, glossario REACH, CLP e ADR.

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