

7. **Apply:** For a strong base, the concentration of hydroxide ions $[\text{OH}^-]$ is estimated to be the same as the concentration of the base. The pH of a strong base is found with the equation $\text{pH} = 14 + \log_{10}[\text{OH}^-]$. Based on their concentrations, find the $[\text{OH}^-]$ and pH of each of the strong bases. Check your answers with the Gizmo.

$[\text{NaOH}] = 0.090 \text{ M}$

$[\text{OH}^-] = 1.05$

pH NaOH = 12.95

$[\text{Ca}(\text{OH})_2] = 0.021 \text{ M}$

$[\text{OH}^-] = 1.68$

pH $\text{Ca}(\text{OH})_2 = 12.32$

*Note: For $\text{Ca}(\text{OH})_2$, the OH^- concentration is double the $\text{Ca}(\text{OH})_2$ concentration because there are two OH^- ions in each $\text{Ca}(\text{OH})_2$ molecule.

Activity C:

Weak acids and bases

Get the Gizmo ready:

- Click **Reset**.
- Select **1.00 M NaOH** for the Burette and **Mystery CH_3COOH** for the Flask.



Introduction: Unlike strong acids and bases, **weak acids** and **weak bases** dissociate relatively little in water. Some ions are formed, but the remaining molecules remain whole. As a result, the pH of a weak acid or base is closer to neutral than the pH of a strong acid or base.

When weak acids or bases react with strong bases or acids, the resulting salts often act as bases or acids themselves, causing the pH at the equivalence point to vary from 7.0. This can impact your choice of indicator.

Question: What happens when weak acids and bases are titrated?

1. **Gather data:** For each indicator given in the Gizmo, what is the pH range over which it changes color?

Bromthymol blue:

6.0-7.6

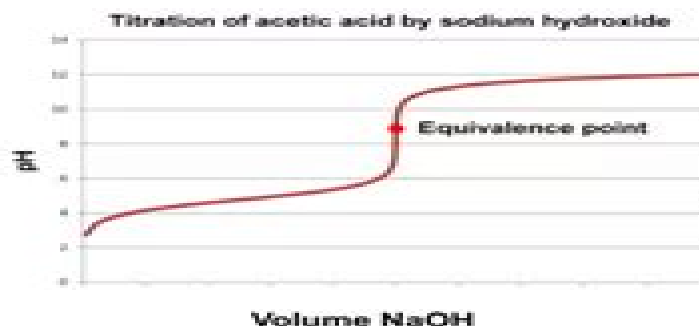
Methyl orange:

3.1-4.4

Phenolphthalein:

8.2

2. **Interpret:** The salt produced by the reaction of acetic acid and sodium hydroxide is a weak base, CH_3COONa . As a result, the pH of the equivalence point is slightly basic. The titration curve for this reaction is shown below:



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