

Multiple Choice: Acid/Base Titrations

1. A solution of known concentration is the definition of a

- A) buffer solution. B) neutral solution. C) standard solution. D) saturated solution

ANS: C

1. A weak acid ($K_a = 4.5 \times 10^{-5}$) is neutralized with a weak base ($K_b = 5.6 \times 10^{-6}$). What is the pH of the resulting solution?

- A) acidic B) basic C) neutral D) pH = 3.2 E) pH = 7

ANS: A

1. The equivalence point of a neutralization reaction is 8.5. What is the best indicator to use?

- A) crystal violet B) bromocresol green C) thymol blue D) methyl red E) alizarin

ANS: C

1. What is the pH of an aqueous solution that is 0.10 M HCOOH ($K_a = 1.8 \times 10^{-4}$) and 0.10 M NaHCO_2 ?

- A) 3.74 B) 10.26 C) 5.62 D) 2.38 E) 5.74

ANS: A

1. Which of the following indicators would be most suitable for the titration of 0.10 M HBr(aq) with 0.10 M KOH(aq) ?

- A) bromothymol blue ($\text{p}K_{\text{in}} = 7.1$) B) thymol blue ($\text{p}K_{\text{in}} = 1.7$)
C) bromophenol blue ($\text{p}K_{\text{in}} = 3.9$) D) alizarin yellow ($\text{p}K_{\text{in}} = 11.2$)
E) methyl orange ($\text{p}K_{\text{in}} = 3.4$)

ANS: A

3. 200.0 mL of 0.200 M HCl is titrated with 0.050 M NaOH . What is the pH after the addition of 100. mL of the NaOH solution?

- A) 0.93 B) 1.45 C) 1.03 D) 0.76 E) 0.82

ANS: A

1. If a 17.0 mL sample of 0.20 M sulfuric acid (H_2SO_4) is titrated with 0.14 M NaOH , it will require how many mL of titrant to reach neutralization?

- A) 9.7 mL B) 48.6 mL C) 389 mL D) 34.0 mL E) 28.6 mL

ANS: B

1. When NaOH(aq) is titrated with HCl(aq) , what is the pH at the equivalence point?

- A. equal to 7 B. greater than 7 C. less than 7 D. undefined

Titration Multiple Choice

Wenbin Ji

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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Practical Book of Pharmaceutical Analysis I Niladry Sekhar Ghosh, The B Pharmacy students are intended to use this laboratory manual The topics covered in the text are generally connected to Pharmaceutical Analysis I The practical work began with an introduction to laboratory techniques used in analysis like weighing and volumetric apparatus and their calibration followed by standardisation procedures as per the syllabus of PCI The assay of specific drugs is then performed in accordance with the syllabus After that electro analytical procedures are provided Appropriate explanations have been given for each practice A few interesting questions for oral exams are provided at the conclusion of each chapter The bibliography is given on the last pages of the book

Analytical Chemistry-A Qualitative and Quantitative Approach Deepak Chowrasia, Nisha Sharma, Book envelops various analytical procedures including their principle and application in chemical and drug analysis

Cracking the AP Chemistry Paul Foglino, Princeton Review (Firm), 2004 The fiercer the competition to get into college the more schools require that students prove themselves in other ways than SAT scores and grade point averages The more expensive college educations become the more students take advantage of the opportunity to test out off first year college courses Includes 2 sample tests with full explanations for all answers The Princeton Review's proven score raising skills and techniques Complete subject review of all the material likely to show up on the AP Chemistry exam

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A Textbook of Pharmaceutical Analysis Dr. Yedlapalli. Govindarao , Mrs. Sakshy Sharma , Dr. Padmanabh B. Deshpande , Mr. Manoj Gangadhar Shinde, Mrs. G. Jahnavi, 2025-05-10 It brings us immense joy to introduce the book Pharmaceutical Analysis This book has been carefully designed to align with the Bachelor of Pharmacy curriculum set by the Pharmacy Council of India We hope it proves valuable to both students and teachers alike We welcome feedback and suggestions on all aspects of the subject and take full responsibility for any inadvertent errors or omissions If any discrepancies are found we would greatly

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