

## Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON\_METALS share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.)  $H_2$  hydrogen is diatomic:



2.)  $F_2$  fluorine is diatomic:



3.)  $O_2$  oxygen is diatomic:



4.)  $N_2$  nitrogen is diatomic. Is this a triple bond?



5.)  $BF_3$  you need 3 fluorine atoms here:



6.) Ammonia  $NH_3$  hint: how many hydrogen atoms are needed?



7.) Carbon dioxide  $CO_2$



8.) Methane  $CH_4$  careful here 4 hydrogen atoms needed



9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!



10.)  $SO_2$  hint: one pair of electrons from sulfur must be slit up for this one to work.



# The Nature Of Covalent Bonding Worksheet Answers

**L Reisser**



## **The Nature Of Covalent Bonding Worksheet Answers:**

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**Chapter Resource 2 Chemistry of Life Biology** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004

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*Merrill Chemistry* Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Glencoe Science** McGraw-Hill Staff,2001-06 **Polar Covalence** R Sanderson,2012-12-02 Polar Covalence provides a detailed account of a successful approach to understanding chemistry from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic properties This text then describes chemical bonding which can only occur when the nuclei of both atoms can attract the same electrons Other chapters consider the bond energy of multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the relative roles of principles and practice in the teaching of inorganic and general

chemistry This book is a valuable resource for chemists and students **Chemical Bonds** Phillip Manning, 2009 Modern life is made up of a mind boggling array of materials A simple drinking cup for example might be made of Styrofoam paper or glass depending on the drinkers needs at the moment Home storage cabinets can be made of metal wood or plastic Space shuttles are assembled from silicon steel and hundreds of other materials All of these items owe their properties to the chemical bonds between the atoms that make up the substance Chemical Bonds examines the nature of the chemical bonds answering fundamental questions about how they form how they are broken and how they help define life as we know it

Nature of the Chemical Bond William A. Goddard (III.), 1980

**The Covalent Bond** Henry Sinclair Pickering, 1977

**Nature of the Chemical Bond** William A. Goddard, 1986

*Chemical Binding and Structure* J. E. Spice, 2016-01-22

Chemical Binding and Structure describes the chemical binding and structure in terms of current chemical theory This book is composed of 13 chapters and starts with a presentation of the principles of the old and modified quantum theory and its application The next chapters cover some basic topics related to chemical binding and structure including electrons the periodic table the electrovalent and covalent bonds and molecular geometry These topics are followed by discussions on the nature of the bond in transition metal complexes electronic and crystal structure crystallinity and other states of matter The concluding chapters are devoted to some analytical techniques for structure determination such as diffraction and spectroscopic methods This book is of value to high school and college chemistry teachers and students

*Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know* Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar

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