

Name: _____

Quiz: Types of Chemical Reactions

True/False (1 point each)

1. In a decomposition reaction, a single compound breaks down into two or more simpler substances.
2. Combustion reactions always involve oxygen as a reactant.
3. A chemical reaction that releases energy is called an endothermic reaction.
4. In a double-replacement reaction, the positive and negative ions of two reactants switch places.
5. The reactants in a chemical equation are found on the right side of the arrow.

Multiple Choice (1 point each)

6. Which of the following is the general form for a synthesis reaction? a) $AB \rightarrow A + B$ b) $A + B \rightarrow AB$ c) $A + BC \rightarrow AC + B$ d) $AB + CD \rightarrow AD + CB$
7. What type of reaction is represented by the following equation: $2H_2O(g) \rightarrow 2H_2(g) + O_2(g)$? a) Synthesis b) Decomposition c) Single-replacement d) Double-replacement
8. In a single-replacement reaction, a more reactive element replaces a less reactive element in a compound. Which of the following metals is the MOST reactive? a) Copper (Cu) b) Zinc (Zn) c) Silver (Ag) d) Gold (Au)
9. Which of the following is an example of a combustion reaction? a) $2Na(s) + Cl_2(g) \rightarrow 2NaCl(s)$ b) $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$ c) $CH_4(g) + 2O_2(g) \rightarrow CO_2(g) + 2H_2O(l)$ d) $Fe(s) + CuSO_4(aq) \rightarrow FeSO_4(aq) + Cu(s)$
10. What type of reaction is represented by the following equation: $AgNO_3(aq) + NaCl(aq) \rightarrow AgCl(s) + NaNO_3(aq)$? a) Synthesis b) Decomposition c) Single-replacement d) Double-replacement
11. Which of the following is NOT a sign that a chemical reaction has occurred? a) Change in color b) Formation of a precipitate c) Change in state d) Production of gas
12. What is the name for a solid that forms from a solution during a chemical reaction? a) Catalyst b) Precipitate c) Reactant d) Product

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