

## Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON\_METALS share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.)  $H_2$  hydrogen is diatomic:



2.)  $F_2$  fluorine is diatomic:



3.)  $O_2$  oxygen is diatomic:



4.)  $N_2$  nitrogen is diatomic. Is this a triple bond?



5.)  $BF_3$  you need 3 fluorine atoms here:



6.) Ammonia  $NH_3$  hint: how many hydrogen atoms are needed?



7.) Carbon dioxide  $CO_2$



8.) Methane  $CH_4$  careful here 4 hydrogen atoms needed



9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!



10.)  $SO_2$  hint: one pair of electrons from sulfur must be slit up for this one to work.



# The Octet Rule In Covalent Bonding Worksheet Answers

**Shirley Soto**



## **The Octet Rule In Covalent Bonding Worksheet Answers:**

**Exploration of the Structure of Atom** Chandan Sengupta, First Publication April 2019 Revised Publication October 2022 Second Revised Edition July 2023 Third Revised Edition December 2024 Published by Chandan Senguta Printed by IECIT Printing and ublication Services Resource Centre Arabinda Nagar Bankura 722101 W B India Topics Covered Atoms and Molecules Structure of Atom Bonding Mechanism and Chemical Reactions Mechanism of Bonding This book is suitable for students of Class 9 to Class 11 Students aspiring for Pre Medical Entrance Examination can also get adequate support Additional Hard Copies can also be obtained from Chandan Sukumar Sengupta Arabinda Nagar Bankura 722101 WB Write to Us for more materials *Holt Chemistry* Ralph Thomas Myers,2004 **Holt Chemistry** Holt Rinehart & Winston,2003-01-24 **Chemical Bonding and the Geometry of Molecules** George E. Ryschkewitsch,1963

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more Molecules and the Chemical Bond Henry A. Bent, 2013-09-30 MOLECULES and the Chemical Bond is about understanding Schrödinger's equation for chemical systems. In his famous Lectures on Physics, Richard Feynman quotes Paul Dirac on what it means to understand an equation: "I understand what an equation means" said Dirac, "if I have a way of figuring out the characteristics of its solutions without actually solving it." That hits the nail on the head. It's precisely what Conceptual Valence Bond Theory does for Schrödinger's equation. A physical understanding of an equation, adds Feynman, is a completely unmathematical, imprecise and inexact thing but absolutely necessary for a physicist. It unfolds in MCB in two stages described by Newton as a stage of Analysis, a union of observations and inductions, and a stage of Synthesis, use of inductions accepted as first principles to explain observations. The book's chief vehicle for creating an intuitive understanding of solutions of Schrödinger's equation is the world's largest and to the author's knowledge, virtually only library of line drawings of exclusive orbital models of chemical species, electron density profiles. By focussing attention on fundamental physical principles and by avoiding use of atomic orbitals and thereby mathematical complexities associated with Schrödinger's equation, the only source of atomic orbitals, the book's essays provide a scientifically sound, student-friendly introduction to modern valence theory. Repetition of fundamental ideas here and there is intended to make individual essays understandable and interesting, each by itself, so that readers may examine them in any order in leisurely walks, so to speak, in the big garden that is valence theory, picking bouquets to their liking.

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