

Chemistry

Time: 2 hours

Marks: 60

- Five isomeric para-disubstituted aromatic compounds A to E with molecular formula $C_8H_8O_2$ were given for identification. Based on the following observations, give structures of the compounds.
 - Both A and B form a silver mirror with Tollen's reagent; also B gives a positive test with $FeCl_3$ solution.
 - C gives positive iodoform test.
 - D is readily extracted in aqueous $NaHCO_3$ solution.
 - E on acid hydrolysis gives 1,4-dihydroxybenzene.[5]
- 500 mL of 0.2 M aqueous solution of acetic acid is mixed with 500 mL of 0.2 M HCl at $25^\circ C$.
 - Calculate the degree of dissociation of acetic acid in the resulting solution and pH of the solution.
 - If 6 g of NaOH is added to the above solution, determine the final pH. [Assume there is no change in volume on mixing; K_a of acetic acid is $1.75 \times 10^{-5} \text{ mol L}^{-1}$].[5]
- Deduce the structures of $[NiCl_4]^{2-}$ and $[Ni(CN)_4]^{2-}$ considering the hybridization of the metal ion. Calculate the magnetic moment (spin only) of the species. [5]
- The density of the vapour of a substance at 1 atm pressure and 500 K is 0.36 kg m^{-3} . The vapour effuses through a small hole at a rate of 1.33 times faster than oxygen under the same condition.
 - Determine (i) molecular weight, (ii) molar volume, (iii) compression factor (Z) of the vapour and (iv) which forces among the gas molecules are dominating, the attractive or the repulsive?
 - If the vapour behaves ideally at 1000 K, determine the average translational kinetic energy of a molecule.[5]
- Write balance equations for the reactions of the following compounds with water;
 - Al_2C_3
 - $CaNCN$
 - BF_3
 - NCl_3
 - XeF_4[5]
- Identify X, Y and Z in the following synthetic scheme and write their structures.
$$CH_3CH_2C \equiv C-H \xrightarrow[(II) CH_3CH_2Br]{(I) NaNH_2} X \xrightarrow{H_2/Pd-BaSO_4} Y \xrightarrow{\text{alkaline } KMnO_4} Z$$
Is the compound Z optically active? Justify your answer. [5]
- How is boron obtained from borax? Give chemical equations with reaction conditions. Write the structure of B_2H_6 and its reaction with HCl.

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