

VALENCE ELECTRONS

Name _____

The valence electrons are the electrons in the outermost principal energy level. They are always "s" or "s and p" electrons. Since the total number of electrons possible in s and p sublevels is eight, there can be no more than eight valence electrons.

Determine the number of valence electrons in the atoms below.

Example: carbon

Electron configuration is $1s^2$ $2s^2 2p^2$.

Carbon has 4 valence electrons.

1. fluorine _____

11. lithium _____

2. phosphorus _____

12. zinc _____

3. calcium _____

13. carbon _____

4. nitrogen _____

14. iodine _____

5. iron _____

15. oxygen _____

6. argon _____

16. barium _____

7. potassium _____

17. aluminum _____

8. helium _____

18. hydrogen _____

9. magnesium _____

19. xenon _____

10. sulfur _____

20. copper _____

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Victor M. Corman



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The purpose of this study is to investigate the place of valence as a topic in high school chemistry with a view to finding possible means of closing the gap between what is actually achieved at the secondary level and what seems to be a desirable point as which to initiate college study Some of the factors responsible for this might well be the confusion in terminology the lag between the enunciation of theories and their appearance in text books the difficulty of symbolic representation of atomic and molecular structure and the closed mind many students develop toward change The extent to which these difficulties can be resolved in the light of research studies historical development of the theory future work to be encountered and the level of maturity of the high school pupil will be considered Historical works surveys of educational research scientific and educational journals text books courses of studies and works on methodology have been consulted The Introduction and Statement of the Problem 1 2 3

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