

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

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Table of Contents Study Guide Praxis Chemistry Printable

1. Understanding the eBook Study Guide Praxis Chemistry Printable
 - The Rise of Digital Reading Study Guide Praxis Chemistry Printable
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Praxis Chemistry Printable
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Praxis Chemistry Printable
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Praxis Chemistry Printable

- Personalized Recommendations
- Study Guide Praxis Chemistry Printable User Reviews and Ratings
- Study Guide Praxis Chemistry Printable and Bestseller Lists
- 5. Accessing Study Guide Praxis Chemistry Printable Free and Paid eBooks
 - Study Guide Praxis Chemistry Printable Public Domain eBooks
 - Study Guide Praxis Chemistry Printable eBook Subscription Services
 - Study Guide Praxis Chemistry Printable Budget-Friendly Options
- 6. Navigating Study Guide Praxis Chemistry Printable eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Praxis Chemistry Printable Compatibility with Devices
 - Study Guide Praxis Chemistry Printable Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Praxis Chemistry Printable
 - Highlighting and Note-Taking Study Guide Praxis Chemistry Printable
 - Interactive Elements Study Guide Praxis Chemistry Printable
- 8. Staying Engaged with Study Guide Praxis Chemistry Printable
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Praxis Chemistry Printable
- 9. Balancing eBooks and Physical Books Study Guide Praxis Chemistry Printable
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Praxis Chemistry Printable
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Praxis Chemistry Printable
 - Setting Reading Goals Study Guide Praxis Chemistry Printable
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Praxis Chemistry Printable

- Fact-Checking eBook Content of Study Guide Praxis Chemistry Printable
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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