## Physical Science 3rd Quarter Exam Reviewer

#### Origin of Elements

Nucleosynthesis- formation of elements through nuclear reactions

#### Lesson 1.1 The Light Elements

- Bigb ang Nucleosynthesis the process where light elements were formed
- Hydrogen and Helium- the first light elements to be formed
- Rapid drop of temperature from 10<sup>12</sup> to 10<sup>6</sup>
- Quarks and Gluons- condensed into nucleons which formed the light elements

#### Lesson 1.2 The Heavy Elements

- Stellar Nucleosynthesis the formation of heavy elements by fusion of lighter nuclei in the interior of stars.
- Nebulas cloud of Hydrogen and Helium
- Protostar mass of H and He
- Stars can only produced elements up to Iron

# Lesson 1.3 Formation of Elements heavier than Iron

- Supernova nucleosynthesis the explosive death of star; the creation of new elements based on the intensity of the burst
- Extreme temperatures and Abundant neutrons - Two key characteristics of Supernova explosion
- Neutron capture when the nuclei captures a electrically neutral neutron to form heavier elements than iron

#### Lesson 1.4 Chemical Elements and Isotopes in the Universe

- Isotopes -is a form of an element that has the same atomic number of the original element but with different atomic mass or mass number.
- Stable Isotope are not radioactive
- Unstable Isotope are radioactive; decays and emits positron and neutrino to form a new element.
- Positron same mass as the electron but its charge is positive
- Neutrino massless and has no charge

#### II. The Atomic Theory

Lesson 2.1 The Atomic Theory: From the Ancient Greeks up to the Present

- Empedocles' Theory proposed that there were four fundamental substances. These substances constitute all matter.
- Democritus' Theory he claimed that everything is made up of atoms.;he called the very tiny pieces of matter "atomos" which means indivisible; the atom is the smallest possible piece of matter.

Aristotles' Theory
accepted the ideas of Empedocles and even
added his own idea that the four substances
could be transformed into other substances,
he associated the four fundamental
elements-earth, water, air and fire with four
qualities: hot, moist, cold and dry.

 Alchemys' Theory of Matter based on Aristotles' theory of two pairs of opposed qualities and the four elements produced by pairwise conjunction of the four qualities.; when these elements interchange qualities, they produce other substance.

#### Lesson 2.2 The Structure of the Atom

- John Dalton proposed an atomic theory about indivisible building blocks of matter called atoms; atoms are spherical in shape and are always in motion.
- Four Main Concepts of Dalton's Atomic Theory

 A. matter is composed of tiny, indivisible particles called atoms.

all atoms of the same element are identical in mass and properties.

compounds are formed by a combination of two or more different kinds of atoms.

 achemical reaction is a rearrangement of atoms.

Scientific Laws About Chemical Reactions

A. 1. Law of Conservation of Mass
Antoine Lavoisier in 1789
This law states that the total mass in a chemical reaction remains constant.

B. Law of Definite Proportions
- Joseph Louis Proust in 1799

# What Is 3rd Exam Of Ss2015 Physical

John Devaney, Caroline Bradbury-Jones, Rebecca J. Macy, Carolina Øverlien, Stephanie Holt

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#### Table of Contents What Is 3rd Exam Of Ss2015 Physical

- 1. Understanding the eBook What Is 3rd Exam Of Ss2015 Physical
  - The Rise of Digital Reading What Is 3rd Exam Of Ss2015 Physical
  - Advantages of eBooks Over Traditional Books
- 2. Identifying What Is 3rd Exam Of Ss2015 Physical
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - ∘ Features to Look for in an What Is 3rd Exam Of Ss2015 Physical
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from What Is 3rd Exam Of Ss2015 Physical
  - Personalized Recommendations
  - What Is 3rd Exam Of Ss2015 Physical User Reviews and Ratings
  - What Is 3rd Exam Of Ss2015 Physical and Bestseller Lists
- 5. Accessing What Is 3rd Exam Of Ss2015 Physical Free and Paid eBooks
  - What Is 3rd Exam Of Ss2015 Physical Public Domain eBooks
  - What Is 3rd Exam Of Ss2015 Physical eBook Subscription Services
  - What Is 3rd Exam Of Ss2015 Physical Budget-Friendly Options
- 6. Navigating What Is 3rd Exam Of Ss2015 Physical eBook Formats

- o ePub, PDF, MOBI, and More
- What Is 3rd Exam Of Ss2015 Physical Compatibility with Devices
- What Is 3rd Exam Of Ss2015 Physical Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of What Is 3rd Exam Of Ss2015 Physical
  - Highlighting and Note-Taking What Is 3rd Exam Of Ss2015 Physical
  - Interactive Elements What Is 3rd Exam Of Ss2015 Physical
- 8. Staying Engaged with What Is 3rd Exam Of Ss2015 Physical
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - $\circ\,$  Following Authors and Publishers What Is 3rd Exam Of Ss2015 Physical
- 9. Balancing eBooks and Physical Books What Is 3rd Exam Of Ss2015 Physical
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection What Is 3rd Exam Of Ss2015 Physical
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine What Is 3rd Exam Of Ss2015 Physical
  - Setting Reading Goals What Is 3rd Exam Of Ss2015 Physical
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of What Is 3rd Exam Of Ss2015 Physical
  - Fact-Checking eBook Content of What Is 3rd Exam Of Ss2015 Physical
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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