

STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and product that allow us to calculate amounts of each

Using Ratios

Reactant and product amounts are related

Each coefficient in the balanced equation can be used to define a ratio between the individual reactants and products of the chemical reaction for the given balanced formula and division of units.

Consider the reaction:

$$2\text{NaCl} + \text{MgSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{MgCl}_2$$

We can write these ratios:

moles of sodium	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$
moles of sulfate	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$
moles of sodium	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$
moles of sodium	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$
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moles of sulfate	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$

Example: Finding the Limiting Reactant in the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

5.37 grams of NaCl react with 5.63 grams of Mg . What is the limiting reactant and what mass of MgSO_4 can be formed?

How many moles of NaCl react? $\frac{5.37\text{g NaCl}}{58.44\text{g/mol}} = 0.0919\text{mol NaCl}$

How many moles of MgSO_4 react? $\frac{5.63\text{g MgSO}_4}{120.36\text{g/mol}} = 0.0468\text{mol MgSO}_4$

How many moles of Na_2SO_4 can be formed? $0.0919\text{mol NaCl} \times \frac{1\text{mol Na}_2\text{SO}_4}{2\text{mol NaCl}} = 0.0459\text{mol Na}_2\text{SO}_4$

How many moles of MgCl_2 can be formed? $0.0468\text{mol MgSO}_4 \times \frac{1\text{mol MgCl}_2}{1\text{mol MgSO}_4} = 0.0468\text{mol MgCl}_2$

5.63 grams of MgSO_4 will form and NaCl is the limiting reactant

Using Mole Ratios

From the last page, we looked at the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

We also defined ratios of moles of each substance using the coefficients of the balanced equation.

How do we use the ratios we defined?

How many moles of Mg can be produced from 4 moles of NaCl in the reaction above?

We choose the ratio of moles of Mg to moles of NaCl resulting in the given number of moles of NaCl . By writing the ratio $\frac{1\text{mol Mg}}{2\text{mol NaCl}}$ in the denominator, we obtain the ratio moles of Mg and are left with moles of NaCl .

$$4\text{ moles NaCl} \times \frac{1\text{mol Mg}}{2\text{mol NaCl}} = 2\text{ moles Mg}$$

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