

- When an oxygen atom forms an oxide ion, O^{2-} , it achieves a noble gas configuration.
 - What is a noble gas configuration?
It fills outer level
 - What noble gas does oxide ion, O^{2-} , have the electron configuration of?
Neon
- When an aluminum atom forms an aluminum ion, Al^{3+} , it becomes more stable.
 - Are electrons lost or gained during this change?
lost
 - How many electrons are lost/gained?
Three electrons are lost.
 - Why is an aluminum ion, Al^{3+} , more stable than an aluminum atom?
 $Al = 1s^2 2s^2 2p^6 3s^2 3p^1$. Al^{3+} has none. It achieves an $1s^2 2s^2 2p^6$ have the same electron arrangement as neon giving it a full outer level which is more stable.
- Why do most transition elements exhibit more than one possible oxidation number?
Because they can lose electrons from different energy levels and do not have to lose them all at once but instead can lose some first and so on.
- Nearly all transition elements can have oxidation numbers of ± 1 . Explain why?
More than 2 valence electrons have electrons which they can lose.
- The electron configuration of iron, Fe, is $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$.
 - Write the electron configuration for Fe^{2+} .
 $1s^2 2s^2 2p^6 3s^2 3p^4$
 - Write the electron configuration for Fe^{3+} .
 $1s^2 2s^2 2p^6 3s^2 3p^4$
 - Explain why Fe^{2+} ion tends to form.
This has a 1/2 full outer shell which is an stable as it can get closer to nearest noble gas electrons and the Fe^{2+} .

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