

## Assessment

**Chapter Test B****Teacher Notes and Answers****8**      *Chemical Equations and Reactions***TEST B**

1. b
2. a
3. b
4. d
5. a
6. b
7. d
8. a
9. coefficient
10. liquid
11. reversible reaction
12.  $A + X \rightarrow AX$
13.  $AX \rightarrow A + X$
14.  $A + BX \rightarrow AX + B$
15.  $AX + BY \rightarrow AY + BX$
16. metal hydroxides
17. combustion reaction
18. electrolysis
19. activity series
20. chemical equation
21. conservation of mass
22. gaseous
23. heated
24. A word equation and a formula equation are both qualitative. Although they both show the reactants and products of a chemical reaction, they give no information about amounts of reactants or products. A chemical equation, however, does show the relative amounts of the reactants and products in a chemical reaction.
25. Three indications of a chemical reaction are the radiation of heat and light, the production of a gas, and the formation of a precipitate.
26. zinc sulfide + oxygen  $\rightarrow$   
zinc oxide + sulfur dioxide
27. sodium iodide + chlorine  $\rightarrow$   
sodium chloride + iodine
28.  $Mg + O_2 \rightarrow MgO$
29.  $CaCO_3 \rightarrow CaO + CO_2$
30. decomposition reaction
31. single-displacement reaction
32. synthesis reaction
33. double-displacement reaction
34.  $2AgNO_3 + CuCl_2 \rightarrow 2AgCl + Cu(NO_3)_2$
35.  $2PbO_2 \rightarrow 2PbO + O_2$
36.  $Zn(OH)_2 + 2CH_3COOH \rightarrow$   
 $Zn(CH_3COO)_2 + 2H_2O$
37.  $Mg(s) + 2H_2O(g) \rightarrow Mg(OH)_2(aq) + H_2(g)$
38. no reaction
39.  $Cl_2(g) + MgBr_2(aq) \rightarrow MgCl_2(aq) + Br_2(l)$
40.  $Zn(s) + 2HCl(aq) \rightarrow ZnCl_2(aq) + H_2(g)$
41. no reaction
42.  $H_2(s) + CuCl_2(aq) \rightarrow H_2Cl_2(aq) + Cu(s)$
43. no reaction
44.  $Mg(s) + Co(NO_3)_2(aq) \rightarrow$   
 $Mg(NO_3)_2(aq) + Co(s)$

# Teaching Resources Chapter 36 Answer Key

**SJ Ball**



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