

- The electron configuration of hydrogen is like that of Group 1 metals.
- Magnesium is more reactive than iron and will replace iron in a compound. If the iron corrodes, the magnesium will replace it, restoring the iron.
- The number of oxygen atoms is not balanced because parentheses were not placed around the hydroxide ion.

#### Applying Scientific Methods

- $2\text{CuS(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{CuO(s)} + 2\text{SO}_2\text{(g)}$   
 $2\text{Cu}_2\text{S(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Cu}_2\text{O(s)} + 2\text{SO}_2\text{(g)}$
- 2, 5, 4, 4, 2, 2
- $\text{CuO(s)} + \text{C(s)} \rightarrow \text{Cu(l)} + \text{CO(g)}$ ;  $\text{Cu}_2\text{O(s)} + \text{C(s)} \rightarrow 2\text{Cu(l)} + \text{CO(g)}$
  - Sample answer:  $\text{Mg(l)} + \text{CuO(s)} \rightarrow \text{Cu(l)} + \text{MgO(s)}$
- $2\text{Fe}_2\text{O}_3\text{(s)} + 3\text{C(s)} \rightarrow 4\text{Fe(l)} + 3\text{CO}_2\text{(g)}$
- $4\text{Fe(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Fe}_2\text{O}_3\text{(s)}$
- synthesis
- $\text{Na(l)} + \text{KCl(l)} \rightarrow \text{NaCl(l)} + \text{K(g)}$
- $\text{Ca(l)} + 2\text{CsCl(l)} \rightarrow \text{CaCl}_2\text{(l)} + 2\text{Cs(g)}$

## CHAPTER 10

### MiniLab 10 – Percent Composition and Gum

#### Analysis

- Mass dissolved = (initial mass) – (final mass) =  $3.11\text{ g} - 2.84\text{ g} = 0.27\text{ g}$  dissolved sweeteners and flavorings
- Mass dissolved = (initial mass) – (final mass) =  $3.11\text{ g} - 2.75\text{ g} = 0.36\text{ g}$  dissolved sweeteners and flavorings
- whole piece of gum:  $0.27\text{ g}/3.11\text{ g} \times 100 = 8.7\%$   
small pieces of gum:  $0.36\text{ g}/3.11\text{ g} \times 100 = 12\%$
- Because more sweetener dissolves when the surface area is larger, the inference is that the sweeteners are mixed throughout the gum.

**Expected Results:** The gum will lose mass after being placed in water because the sweeteners are

water-soluble. The cut piece with the larger surface area will expose more of the soluble chemicals to the water, and will have a greater loss of mass.

### ChemLab 10 – Hydrated Crystals

#### Pre-Lab

- To obtain the mass of water, subtract the mass of the crucible, cover, and anhydrous  $\text{MgSO}_4$  from the mass of the crucible, cover, and hydrated  $\text{MgSO}_4$ . To obtain the mass of anhydrous  $\text{MgSO}_4$ , subtract the mass of the crucible and cover from the mass of the crucible, cover, and anhydrous  $\text{MgSO}_4$ .
- Divide by the molar mass.
- Calculate the ratio of moles of water to moles of anhydrous  $\text{MgSO}_4$ .

| Mass Data and Observations of Epsom Salts            |                               |
|--|-------------------------------|
| Observations of hydrated $\text{MgSO}_4$             | Shiny, translucent crystals   |
| Mass of crucible and lid                             | 11.36 g                       |
| Mass of crucible, lid, and hydrated $\text{MgSO}_4$  | 14.36 g                       |
| Mass of hydrated $\text{MgSO}_4$                     | 3.00 g                        |
| Mass of crucible, lid, and anhydrous $\text{MgSO}_4$ | 12.83 g                       |
| Mass of anhydrous $\text{MgSO}_4$                    | 1.47 g                        |
| Mass of water in hydrated $\text{MgSO}_4$            | 1.53 g                        |
| Moles of anhydrous $\text{MgSO}_4$                   | 0.0122 mol                    |
| Moles of water in hydrated $\text{MgSO}_4$           | 0.0849 mol                    |
| Observation of anhydrous $\text{MgSO}_4$             | Bright white, opaque, powdery |

#### Analyze and Conclude

- $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$
- The hydrated  $\text{MgSO}_4$  is shiny and translucent. The anhydrous  $\text{MgSO}_4$  is opaque and bright white.

# Teaching Transparency Worksheets Chemistry Answers

## Chapter 1

**Daniela Niemeyer**



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