



YAMAHA

BRUIN 350
2004-2006

GRIZZLY 350
2007-2008
Service Manual



LIT-11616-BN-35

Yamaha Bruin 350 Service Manual

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web an solution initially contains a catalytic amount of an enzyme with $k_m = 1.5 \text{ mM}$ 0.25 M of substrate and no product after 45 seconds the solution contains $25 \text{ }\mu\text{M}$ of product find v_{max} and the concentration of product after 20 minutes hint $s = k_m$
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web a reaction and its experimental rate law are represented below $2 \text{ ICl(g)} + \text{H}_2\text{(g)} \rightarrow 2 \text{ HCl(g)} + 2 \text{ ICl}_2\text{(g)}$ rate $k [\text{ICl}]^2 [\text{H}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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web in order to have successful collisions the colliding particles must have both the proper amount of energy and the proper fill in blank draw an energy diagram which shows the shape of the curve you might expect for the reaction in this question the overall reaction is exothermic

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textbook if you get stuck try asking another group for help 1 for the reaction below if substance a is disappearing at a rate of $1.82 \times 10^{-2} \text{ mol l}^{-1} \text{ s}^{-1}$ at what rate is c appearing $3 \text{ a} + 3 \text{ b} \rightarrow 5 \text{ c} + 2 \text{ d}$

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web reaction rates google classroom $\text{mg} + \text{s} \rightarrow 2 \text{ hcl} + \text{a} + \text{q}$ $\text{mgcl}_2 + 2 \text{ a} + \text{q} \rightarrow \text{h}_2 + \text{g}$ the reaction between $\text{mg} + \text{s}$ and $\text{hcl} + \text{a} + \text{q}$ is represented by the equation above in a kinetics experiment a 0.080 g sample of $\text{mg} + \text{s}$ ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M $\text{hcl} + \text{a} + \text{q}$ at 25 °C

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