

$$\frac{PV}{P} = \frac{nRT}{P}$$

$$V = \frac{nRT}{P}$$

$$V = 8.28 \text{ L}$$

$$R = .0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

$$T = 273 + 32 = 305 \text{ K}$$

$$P = 1 \text{ atm}$$

$$V = ? \text{ L}$$

$$n = .343 \text{ mol}$$

Pod 10 7:58 PM

SkillPractice.pdf

Skill Practice 36

## Gases and Moles Practice

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_

- Under water where the temperature is  $17^{\circ}\text{C}$  and the pressure is  $394 \text{ kPa}$ , a diver inhales  $2.1 \text{ L}$  of air from his SCUBA tank.
  - How many moles of gas are in his lungs?
- If the diver returns to the surface without exhaling where the temperature is  $32^{\circ}\text{C}$  and the pressure changes to  $100.2 \text{ kPa}$ , what will the volume of the air in his lungs be?

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what 33 European countries have been doing in order to raise the skills levels of low skilled workers address the problem of the high number of early school leavers combat social exclusion ensure the efficiency effectiveness quality of adult learning

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