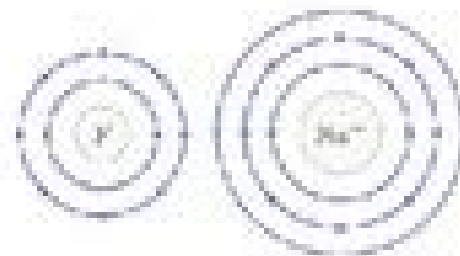


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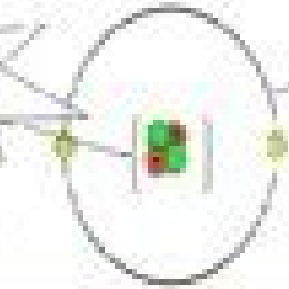
Chemistry is identification of the substances of which **matter** is composed; the investigation of their properties and the ways in which they **interact, combine, and change**; and the use of these processes to form **new substances**.

Lewis Dot Structure diagrams that represent the valence electrons of an atom:



Some elements have more than one shell of electrons. The shells (made of discrete orbits) around the core.

The center is the **NUCLEUS** which are **POSITIVE** and **NEUTRONS** which are **NEUTRAL**.



Outside the nucleus there are **ELECTRON SHELLS** or **ENERGY LEVELS** which hold most **NEUTRALLY** charged particles are called **ELECTRONS**.

Periodic Table:

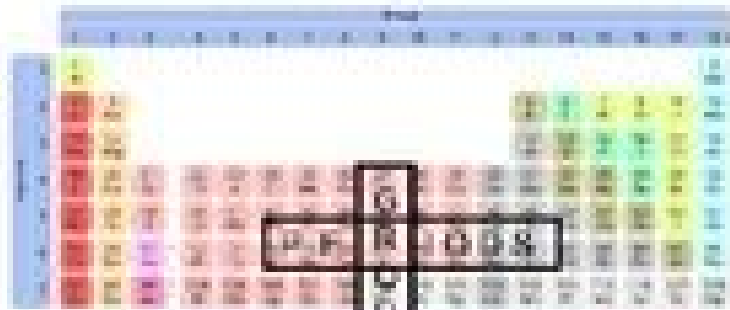
Metals: Alkali Earth Metals, Transition Metals, Alkali Metals, Lanthanides, and Actinides

Non-Metals: Halogens, Noble Gases

Types of Elements:

Metals: lustrous, ductile, malleable, and good conductors of electricity

Non-metals: brittle (for solid elements), lack luster, and are insulators



P, F, Br, O, S
 Halogens
 A, P, E, N, A, N

Group 17 Elements have 7 **VALENT** **ELECTRONS** which are highly **REACTIVE**.

Periods: Shows how many shells or energy levels each element has for its electrons.

A, P, E, N, A, N

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Judd E. Hollander



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