

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kJoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kJoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

Skill Sheet Periodic Trends

Clifford Lane

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

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Table of Contents Skill Sheet Periodic Trends

1. Understanding the eBook Skill Sheet Periodic Trends
 - The Rise of Digital Reading Skill Sheet Periodic Trends
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Sheet Periodic Trends
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Sheet Periodic Trends
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Sheet Periodic Trends
 - Personalized Recommendations
 - Skill Sheet Periodic Trends User Reviews and Ratings
 - Skill Sheet Periodic Trends and Bestseller Lists
5. Accessing Skill Sheet Periodic Trends Free and Paid eBooks
 - Skill Sheet Periodic Trends Public Domain eBooks
 - Skill Sheet Periodic Trends eBook Subscription Services
 - Skill Sheet Periodic Trends Budget-Friendly Options

6. Navigating Skill Sheet Periodic Trends eBook Formats
 - ePub, PDF, MOBI, and More
 - Skill Sheet Periodic Trends Compatibility with Devices
 - Skill Sheet Periodic Trends Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Sheet Periodic Trends
 - Highlighting and Note-Taking Skill Sheet Periodic Trends
 - Interactive Elements Skill Sheet Periodic Trends
8. Staying Engaged with Skill Sheet Periodic Trends
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Sheet Periodic Trends
9. Balancing eBooks and Physical Books Skill Sheet Periodic Trends
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Sheet Periodic Trends
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Skill Sheet Periodic Trends
 - Setting Reading Goals Skill Sheet Periodic Trends
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Skill Sheet Periodic Trends
 - Fact-Checking eBook Content of Skill Sheet Periodic Trends
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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