



Skeleton Equations



Skeleton Equations are equations that are unbalanced, meaning they do not have an equal amount of each element on both sides of the equation



Symbols to indicate state of compound:

(s) – solid (g) – gas (l) – liquid
(aq) – aqueous (dissolved in water)

Any diatomic element found by itself must have a subscript of 2!

Task #1: Convert the following word equations to skeleton equations. Be sure to include the state of matter that each compound has as a subscript. Also, state the type of reaction that takes place.

Example: Silver nitrate + copper → silver + copper nitrate

Solution: $\text{AgNO}_3 \text{ (aq)} + \text{Cu (s)} \rightarrow \text{Ag (s)} + \text{Cu(NO}_3)_2 \text{ (aq)}$ This is a single displacement reaction

(a) Iron + sulphur → Iron(II) sulphide

(b) Hydrogen gas + chlorine gas → hydrogen chloride

(c) Magnesium + oxygen gas → magnesium oxide

(d) Aluminum oxide → aluminum + oxygen gas

(e) Hydrogen gas + oxygen gas → water

(f) Hydrochloric acid + sodium hydroxide → sodium chloride + water

(g) Methane gas + oxygen gas → carbon dioxide + water

(h) Potassium hydroxide + hydrogen bromide → potassium bromide + water

(i) Carbon + oxygen gas → carbon dioxide

(j) Tin (II) oxide + hydrogen gas → tin + water

(k) Write your own word equation, then write the skeleton equation for it

(l) Write your own word equation, then write the skeleton equation for it

Skeletal Equation Worksheet

Kathleen Armour

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Skeletal Equation Worksheet:

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