-4NaOH	$+2F_{2}$	$\rightarrow 4 NaF$	$+2H_{2}O +$	O_2
ii) (conc)				

Reaction with Cl₂, Br₂ and I₂:

$$2NaOH + X_2 \rightarrow NaX + NaOX + H_2O$$

0 (cold / dilute)

ii)
$$6NaOH + 3X_2 \rightarrow 5NaX + NaXO_1 + 3H_2O$$

Sodium carbonate (Na₂CO₃.10H₂O) called washing soda.

Anhydrous sodium carbonate is called soda ash.

Sodium carbonate is prepared by

(i) Le-Blanc process (ii) Ammonia soda process or Solvay process (iii) Electrolytic process

- Le-Blanc process: The raw materials used are NaCl. cone. H₂SO₄, lime stone
- Na₂CO₃ + CaS is called as black ash.
- Solvays process: The raw materials used are NaCl. NH₃ and CaCO₃.

The by product in this process is $CaCl_2$.

- Sodium bi carbonate (NaHCO₃) is called as baking soda.
- Na₂CO₃ + H₂O + CO₂ → 2NaHCO₃
- Aqueous solution of NaHCO₃ is alkaline due to anionic hydrolysis (HCO₃).
- Aqueous solution of NaHCO₃ cannot give any colour with phenolphthalein.
- The oxides of Be, Mg, Ca, Sr and Ba are basic therefore these elements are called alkaline earth metals.
- Radium is radio active element.

- The outer electronic configuration is _{RS}². They exhibit +2 oxidation state.
- The most abundant alkaline earth metal is calcium.
- Be and Mg do not give any flame colouration.
- Be shows diagonal relationship with A1.
- Alkaline earth metals form binary compounds with hydrogen called hydrides.

The ionic nature of hydrides increases in the order

$$BeH_2 < MgH_2 < CaH_2 < SrH_2 < BaH_2$$
 (Ionic)

- Alkaline earth metals react with halogens to form their respective halides.
- Except Berylium halides, all the other halides of alkaline earth metals are ionic.
- Alkaline earth metal fluorides are generally insoluble in water except BeF₂.
- Mg is used in the manufacture of alloys:
 - (i) Magnalium: A1 + Mg (85 - 90%) (1 - 15%)
 - (ii) Elektron: Mg + Zn (95%) (5%)
- Gypsum on heating to 120-130⁰C gives plaster of paris. (2CaSO₄.H₂O).
- Plaster of paris sets to hard mass in two stages

$$CaSO_4\frac{1}{2}H_2O + \frac{3}{2}H_2O \xrightarrow{\text{sering inspir}} + CaSO_4 2H_2O$$

 $CaSO_4 \ 2H_2O$ hardware may $\rightarrow CaSO_4 \ 2H_2O$ (ortho rhombic di hydrate) (Monoclinic dihydrate)

- Lime mortar is a mixture of 1 part of slaked lime, 3 parts of sand and water.
- Mortar mixed with cement is called cement mortar.
- Lime stone and clay when heated together give hydraulic mortar. This is used for bleaching purposes and as an antiseptic.

World Of Chemistry 2 1 Answers

Charalampos Karagiannidis, Panagiotis Politis, Ilias Karasavvidis

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