

Mole Concept

Mole concept:

Mole is the counting unit for atoms, molecules and ions and is equal to 6.022×10^{23} . This number is called Avogadro's number and is denoted by N_A .

One mole is the amount of substance that contains as many particles (atoms, molecules or ions) as in 12g of C-12.

Molar mass:

It is the mass of 1 mole of any substance expressed in grams.

In case of atoms it is equal to atomic mass expressed in grams.

In case of molecules it is equal to molecular mass expressed in grams.

1 Mole of
particles
(atoms,
molecules or
ions)

Contains 6.022×10^{23} particles
(atoms,
molecules or
ions)

Mass is
equal to
molar mass
(Atomic mass
or molecular
mass)

Volume is
equal to
molar
volume =
22.4 lit

For example,

1 mole of Oxygen gas	Contains 6.022×10^{23} Oxygen molecules.	32 g	22.4 lit
1 mole of Oxygen atoms	Contains 6.022×10^{23} Oxygen atoms.	16g	22.4 lit
1 mole of Carbon dioxide gas	Contains 6.022×10^{23} Carbon dioxide molecules.	44g	22.4 lit
1 mole of Ammonia gas	Contains 6.022×10^{23} Ammonia molecules.	17g	22.4 lit

Formula

1. No. of moles = given mass / molar mass ($n = m/M$)
2. No. of moles = given no. of particles / Avogadro no. ($n = N/N_A$)
3. Molar mass = given mass / no. of moles ($M = m/n$)
4. Mass of the substance = molar mass \times no. of moles ($m = M \times n$)
5. No. of particles = no. of moles \times Avogadro's no. (no. of particles = $n \times N_A$)
= (Given Mass/ Molar mass) $\times N_A$

The Mole Concept Review Key

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Daily Graphic Elvis Aryeh, 2002-05-07 *Chemistry*³

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