1 mole = 6.02×10^{23} atoms 1 mole = atomic mass (g)



2. How many moles of barrum are in a cample containing 4.25 x 10" atoms of barrum?

Two Step Problems:

RACTICE PROBLEMS:

1. What is the mass of 2.25 x 10²⁵ atoms of lead?

1. How many motes are equal to 625g of crapper?

Convert 2:38 moles of carbon to alloms

4. How many atoms are in 4 Sg of pollassium?

2.25 x 1028 atoms of lead	1 mole	207.28	-	7744.19g = 7740g
	6.02 x 10 ²³ atoms	1 mode		

2. How many atoms are in 10.0g of gold?

10 g gold	1 mole	6.02×10^{23} atoms	-	3.06×10^{22} atoms
	197.0e	1 mole		

The Mole Chapter Worksheet

Ismail Tosun

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