

# Mole Conversions Practice Answer Key

Molar Mass – Calculate the Molar Mass for the following.

- What is the molar mass of Magnesium metal?

$$24.305 \text{ g/mol}$$

- Calculate the molar mass of NaCl

$$\begin{aligned} \text{Na} &: 22.990 \times 1 = 22.990 \\ \text{Cl} &: 35.453 \times 1 = 35.453 \end{aligned} \Rightarrow 58.443 \text{ g/mol}$$

- What is the molar mass of NaOH?

$$\begin{aligned} \text{Na} &: 1 \times 22.990 = 22.990 \\ \text{O} &: 1 \times 15.999 = 15.999 \\ \text{H} &: 1 \times 1.008 = 1.008 \end{aligned} \Rightarrow 39.997 \text{ g/mol}$$

Conversions – Calculate the following the conversions. Use your Mole Map to help you!

- Determine the number of atoms in 0.50 moles He.

$$0.5 \text{ mol He} \times \frac{6.02 \times 10^{23} \text{ atoms He}}{1 \text{ mol He}} = 3.01 \times 10^{23} \text{ atoms He}$$

- Determine the number of moles in 80.00 g O<sub>2</sub>

$$0.2 \times 15.999 = 31.998 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol} = 31.998 \text{ g}$$

- Determine the volume of 0.24 mol of NO<sub>2</sub>.

$$0.24 \text{ mol NO}_2 \times \frac{22.4 \text{ L NO}_2}{1 \text{ mol NO}_2} = 5.376 \text{ L NO}_2$$

$$80.00 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 25.00 \text{ mol O}_2$$

- Determine the number of moles in  $7.5 \times 10^{19}$  formula units of KNO<sub>3</sub>

$$7.5 \times 10^{19} \text{ f.u. KNO}_3 \times \frac{1 \text{ mol KNO}_3}{6.02 \times 10^{23} \text{ f.u. KNO}_3} = 1.2458 \times 10^{-4} \text{ f.u. KNO}_3$$

- Determine the number of moles in 202.22g KNO<sub>3</sub>

$$\begin{aligned} \text{K} &: 1 \times 39.098 = 39.098 \\ \text{N} &: 1 \times 14.007 = 14.007 \\ \text{O} &: 3 \times 15.999 = 47.997 \end{aligned} \Rightarrow 101.102 \text{ g/mol} \Rightarrow 1 \text{ mol} = 101.102 \text{ g}$$

- Determine the number of moles in 32.5L of N<sub>2</sub> at STP

$$32.5 \text{ L N}_2 \times \frac{1 \text{ mol N}_2}{22.4 \text{ L N}_2} = 1.451 \text{ mol N}_2$$

$$202.22 \text{ g KNO}_3 \times \frac{1 \text{ mol KNO}_3}{101.102 \text{ g KNO}_3} = 2.00 \text{ mol KNO}_3$$

- Determine the number of moles in  $3.4 \times 10^{26}$  formula units of NaCl.

$$3.4 \times 10^{26} \text{ f.u. NaCl} \times \frac{1 \text{ mol NaCl}}{6.02 \times 10^{23} \text{ f.u. NaCl}} = 569.784 \text{ f.u. NaCl}$$

- Determine the moles in 36.5L of carbon monoxide

$$36.5 \text{ L CO} \times \frac{1 \text{ mol CO}}{22.4 \text{ L CO}} = 1.629 \text{ mol CO}$$

# The Mole Chapter Answer Key

**R Bogdan**



## **The Mole Chapter Answer Key:**

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