

Worksheet: Mole Calculations & Percent Composition

A. Mole Calculations. You will need a periodic table to look up molecular weights. Do each type of calculation until you feel that you could do it without looking at notes. If you are very comfortable with a problem, skip down to harder ones. The answers will be on the web.

1. _____ How many moles of Na are in 42 g of Na?
2. _____ How many moles of O are in 8.25 g of O?
3. _____ How much does 2.18 mol of Cu weigh?
4. _____ What is the mass of 0.28 mol of iron?
5. _____ How many atoms are in 7.2 mol of chlorine?
6. _____ How many atoms are in 36 g of bromine?
7. _____ How many moles are in 1.0×10^7 atoms?
8. _____ What is the mass of 1.20×10^{25} atoms of sulfur?
9. _____ How many moles of CO molecules are in 52 g of CO?
10. _____ How many moles of C_2H_6 are in 124 g?
11. _____ How many moles of CCl_4 are there in 56 g?
12. _____ How much does 2.50 mol of H_2SO_4 weigh?
13. _____ How much does 0.25 mol of Fe_2O_3 weigh?
14. _____ How many molecules are there in 52 g of CO?
15. _____ How many formula units are in 22.4 g SnO_2 ?
16. _____ How many molecules are in 116 g CCl_4 ?
17. _____ What is the mass of 3.01×10^{23} formula units of Fe_2O_3 ?
18. _____ What is the mass of 1.2×10^{25} molecules of CO?
19. _____ How many O atoms are in 1.25 mol of SO_2 ?
20. _____ How many moles of O atoms do you have when you have 1.20×10^{25} N_2O_5 molecules?
21. _____ How many formula units are in 5.33 mol of $CuCl_2$?
22. _____ How many copper atoms are in 5.33 mol of $CuCl_2$?
23. _____ How many moles of Cl atoms are in 5.33 mol of $CuCl_2$?

The Mole The Central Unit Of Chemistry Worksheet

**Frank Schaffer Publications,Joan
DiStasio**



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