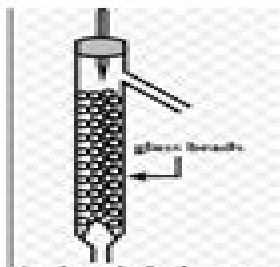


Suggest why chloroethanoic acid is stronger than ethanoic acid	(this question is already discussed a little further down.)		
fill in rate table	1.00 0.16 1.35 I beg someone write what the values were if they remember	K=250 Yes i agree - needs to be 1.4 as all values were to two sig figs Weren't we they to 3 dp though? The rate of the first one was 1.00 so I'm guessing 3sf of 2dp	3
empirical formula	C6H10O3S2	Percentages dudbt add up to 100 fuck me Didn't even see the sulphur, fml same same i thought i meant i was in the fuel area. YAY i got this!!	3
the labeling distillation	for the tube it was fractionating column or glass beads the bottom of the condenser was water in and top was water out	 <p>reject anti bumping granules might accept silica BEADS</p>	3
Draw a line for the reaction with chloroethanoic acid  Suggest why chloroethanoic acid is a stronger acid than ethanoic acid.	<p>Was the line of chloroethanoic acid to the left of the first line, but levelled off at the same point? Yes but it levelled out before, same height though. Ah ok.</p> <p>Dissociates more Chlorine is more electronegative than hydrogen so will withdraw electron density from the carboxyl group meaning that the O-H bond breaks more easily dissociating into H<sup>+</sup> making the solution more acidic</p>	Cl is electron withdrawing. it polarises the bond, weakens the O-H bond which releases more H <sup>+</sup> than ethanoic acid Neg inductive effect? Chloroethanoic acid forms a more stable conjugate base?	3
NMR Question?	<p>Use the integration values and the ratios to identify the structure at two peaks.</p> <p>Draw this structure.</p>	Was there a typo though? Regarding what? No dont think so, it was just because 2 of the peaks were the same environment almost so people got thrown off	5 2 1 2

# Unofficial Mark Scheme Chemistry 2013 Unit 2

**Iseb Iseb**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

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1. The book is structured into several chapters, namely:
    - Chapter 1: Introduction to Unofficial Mark Scheme Chemistry 2013 Unit 2
    - Chapter 2: Essential Elements of Unofficial Mark Scheme Chemistry 2013 Unit 2
    - Chapter 3: Unofficial Mark Scheme Chemistry 2013 Unit 2 in Everyday Life
    - Chapter 4: Unofficial Mark Scheme Chemistry 2013 Unit 2 in Specific Contexts
    - Chapter 5: Conclusion
  2. In chapter 1, this book will provide an overview of Unofficial Mark Scheme Chemistry 2013 Unit 2. The first chapter will explore what Unofficial Mark Scheme Chemistry 2013 Unit 2 is, why Unofficial Mark Scheme Chemistry 2013 Unit 2 is vital, and how to effectively learn about Unofficial Mark Scheme Chemistry 2013 Unit 2.
  3. In chapter 2, the author will delve into the foundational concepts of Unofficial Mark Scheme Chemistry 2013 Unit 2. The second chapter will elucidate the essential principles that need to be understood to grasp Unofficial Mark Scheme Chemistry 2013 Unit 2 in its entirety.
  4. In chapter 3, the author will examine the practical applications of Unofficial Mark Scheme Chemistry 2013 Unit 2 in daily life. This chapter will showcase real-world examples of how Unofficial Mark Scheme Chemistry 2013 Unit 2 can be effectively utilized in everyday scenarios.
  5. In chapter 4, this book will scrutinize the relevance of Unofficial Mark Scheme Chemistry 2013 Unit 2 in specific contexts. This chapter will explore how Unofficial Mark Scheme Chemistry 2013 Unit 2 is applied in specialized fields, such as education, business, and technology.
  6. In chapter 5, the author will draw a conclusion about Unofficial Mark Scheme Chemistry 2013 Unit 2. The final chapter will summarize the key points that have been discussed throughout the book.
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