Enthalpy change for the thermal decomposition of potassium hydrogen carbonate = 5.38 KJ/mol + 0.32 KJ/mol = 5.7 KJ/mol

Enthalpy change for the thermal decomposition of potassium hydrogen carbonate = 5.7 KJ/mol

4. When KHCO₃ is heated, it decomposes according to the following chemical reaction:

$$2 \text{ KHCO}_3(s) \rightarrow \text{K}_2\text{CO}_3(s) + \text{H}_2\text{O}(l) + \text{CO}_2(g) \qquad \Delta H_{\text{rxn}} = ?$$

Apply Hess' Law and use the results from questions 2b and 3b to calculate the enthalpy change for the thermal decomposition of potassium hydrogen carbonate.

Standard Enthalpy Change Of Potassium Hydrogen Carbonate

Robert A. Friedel, H. L. Retcofsky, United States. Bureau of Mines, J. A. Queiser

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