

ANSWERS – THE STATES OF MATTER

WORKSHEET

1. (a) condensation (b) sublimation (c) freezing (d) sublimation (e) melting
(f) sublimation (g) sublimation
2. (a) **SOLIDS MELT WHEN HEATED.** The heat causes the particles to move faster and further apart, breaking the particle attractions. The substance then changes from the solid to the liquid state.

(b) **CHANGES OF STATE ARE PHYSICAL CHANGES** – The particles in the substance simply spread apart or move closer together. No new particles are created.

(c) **GASES DIFFUSE (SPREAD OUT) READILY** – gas particles move very quickly because the attractive forces are negligible. Thus they are able to diffuse (spread out) readily.

(d) **YOUR SKIN FEELS COOL WHEN A LIQUID EVAPORATES FROM IT** – The particles in the liquid absorb heat energy from your skin when they evaporate into the gaseous state. Thus your skin feels cool.

(e) **THE DENSITIES OF GASES ARE MUCH LOWER THAN THOSE OF LIQUIDS AND SOLIDS** – Gases are less dense because their particles are spaced much farther apart. Therefore, there are less particles (less mass) per unit volume.

(f) **GASES ARE MORE COMPRESSIBLE THAN LIQUIDS** – Gases are more compressible than liquids because the spaces between the particles are much larger.

TEXTBOOK

Pg 423

4. (a) when a metal is heated, the molecules gain kinetic energy, vibrational motion increases, and the molecules move further apart. When it is cooled, the reverse happens.

(b) Gases have no fixed volume because they have enough kinetic energy to be relatively unaffected by gravity. Instead, their volume is affected by the size of their container.

(c) Molecules in a gas are spaced farther apart than are molecules in a liquid.
5. Since molecules in gases are little affected by attractions to other molecules, they move much more randomly than molecules in liquids or solids.
6. Molecules in gases have vibrational, rotational, and translational motion.
7. Adding heat increases the kinetic energy of the molecules in a liquid, which increases the vibrational and rotational motion of the molecules, resulting in an increase in volume. With enough energy, they can break the intermolecular attractions and become a gas.

States Of Matter And The Kinetic Theory Michigan Answers

Joacim Rocklöv



States Of Matter And The Kinetic Theory Michigan Answers:

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Table of Contents States Of Matter And The Kinetic Theory Michigan Answers

1. Understanding the eBook States Of Matter And The Kinetic Theory Michigan Answers
 - The Rise of Digital Reading States Of Matter And The Kinetic Theory Michigan Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying States Of Matter And The Kinetic Theory Michigan Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an States Of Matter And The Kinetic Theory Michigan Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from States Of Matter And The Kinetic Theory Michigan Answers
 - Personalized Recommendations
 - States Of Matter And The Kinetic Theory Michigan Answers User Reviews and Ratings
 - States Of Matter And The Kinetic Theory Michigan Answers and Bestseller Lists

5. Accessing States Of Matter And The Kinetic Theory Michigan Answers Free and Paid eBooks
 - States Of Matter And The Kinetic Theory Michigan Answers Public Domain eBooks
 - States Of Matter And The Kinetic Theory Michigan Answers eBook Subscription Services
 - States Of Matter And The Kinetic Theory Michigan Answers Budget-Friendly Options
6. Navigating States Of Matter And The Kinetic Theory Michigan Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - States Of Matter And The Kinetic Theory Michigan Answers Compatibility with Devices
 - States Of Matter And The Kinetic Theory Michigan Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of States Of Matter And The Kinetic Theory Michigan Answers
 - Highlighting and Note-Taking States Of Matter And The Kinetic Theory Michigan Answers
 - Interactive Elements States Of Matter And The Kinetic Theory Michigan Answers
8. Staying Engaged with States Of Matter And The Kinetic Theory Michigan Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers States Of Matter And The Kinetic Theory Michigan Answers
9. Balancing eBooks and Physical Books States Of Matter And The Kinetic Theory Michigan Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection States Of Matter And The Kinetic Theory Michigan Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine States Of Matter And The Kinetic Theory Michigan Answers
 - Setting Reading Goals States Of Matter And The Kinetic Theory Michigan Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of States Of Matter And The Kinetic Theory Michigan Answers
 - Fact-Checking eBook Content of States Of Matter And The Kinetic Theory Michigan Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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