

## Chemical Equilibrium

### Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when
- A. the rate of the forward reaction equals that of the backward reaction.
  - B. the concentration of reactants and products are equal
  - C. the activation energies of the forward and backward reactions are equal
  - D. the reaction involves no enthalpy change.
2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

- A. high temperature, low pressure
  - B. high temperature, high pressure
  - C. low temperature, high pressure
  - D. low temperature, low pressure.
3.  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
- In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because
- A. the ammonia is constantly removed
  - B. an iron catalyst is used
  - C. high pressure is used
  - D. the temperature is held at above  $500^\circ\text{C}$ .

4. The equation refers to the preparation of methanol from synthesis gas.
- $$\text{CO}(\text{g}) + 2\text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$$
- $$\Delta H = -91 \text{ kJ mol}^{-1}$$

The formation of methanol is favoured by

- A. high pressure and low temperature
  - B. high pressure and high temperature
  - C. low pressure and low temperature
  - D. low pressure and high temperature.
5.  $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$
- $$\Delta H = -46 \text{ kJ mol}^{-1}$$

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

	Temperature	Pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

6. When a reversible chemical reaction is at equilibrium,
- A. the concentration of reactants and products remain equal
  - B. the forward reaction is unable to continue
  - C. the concentration of reactants and products remain constant
  - D. the forward and reverse reactions proceed at different rates.

# Static Equilibrium Multiple Choice Questions

**Oswaal Editorial Board**



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## **Table of Contents Static Equilibrium Multiple Choice Questions**

1. Understanding the eBook Static Equilibrium Multiple Choice Questions
  - The Rise of Digital Reading Static Equilibrium Multiple Choice Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Static Equilibrium Multiple Choice Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Static Equilibrium Multiple Choice Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Static Equilibrium Multiple Choice Questions
  - Personalized Recommendations
  - Static Equilibrium Multiple Choice Questions User Reviews and Ratings
  - Static Equilibrium Multiple Choice Questions and Bestseller Lists

5. Accessing Static Equilibrium Multiple Choice Questions Free and Paid eBooks
  - Static Equilibrium Multiple Choice Questions Public Domain eBooks
  - Static Equilibrium Multiple Choice Questions eBook Subscription Services
  - Static Equilibrium Multiple Choice Questions Budget-Friendly Options
6. Navigating Static Equilibrium Multiple Choice Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Static Equilibrium Multiple Choice Questions Compatibility with Devices
  - Static Equilibrium Multiple Choice Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Static Equilibrium Multiple Choice Questions
  - Highlighting and Note-Taking Static Equilibrium Multiple Choice Questions
  - Interactive Elements Static Equilibrium Multiple Choice Questions
8. Staying Engaged with Static Equilibrium Multiple Choice Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Static Equilibrium Multiple Choice Questions
9. Balancing eBooks and Physical Books Static Equilibrium Multiple Choice Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Static Equilibrium Multiple Choice Questions
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Static Equilibrium Multiple Choice Questions
  - Setting Reading Goals Static Equilibrium Multiple Choice Questions
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Static Equilibrium Multiple Choice Questions
  - Fact-Checking eBook Content of Static Equilibrium Multiple Choice Questions
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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