

CHAPTER 14: COVALENT BONDING: ORBITALS

PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- O_2 would be predicted to be diamagnetic, but it is **paramagnetic**

Paramagnetic = has unpaired electrons.

- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as B_2H_6 (electron deficient molecules)



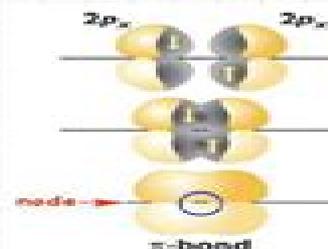
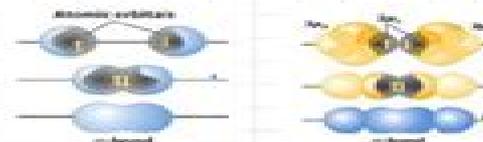
The Lewis structure of O_2 has an O=O double bond with **no unpaired electrons**, i.e. diamagnetic.

VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

TYPES OF OVERLAP: σ and π Bonds

- **Bonding in H_2**
 - Sausage shaped orbital, bond called σ -orbital
 - Called σ orbital because it has **no node** and is **cylindrically symmetrical**
 - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- Bonding in N_2
 - σ orbital is obtained by overlap of **two p_z -orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner
- The resulting orbital has two lobes of electron density and is called a π -bond
- Called π -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- π -bond has no electron density between the nuclei along the bond axis.



University Chemistry Chapter Covalent Bonding Orbitals Study Guide

Victor M. S. Gil



University Chemistry Chapter Covalent Bonding Orbitals Study Guide:

Decoding **University Chemistry Chapter Covalent Bonding Orbitals Study Guide**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**University Chemistry Chapter Covalent Bonding Orbitals Study Guide**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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