

# Experiment 2 - Gravimetric Analysis of a Chloride Salt (G)

## Purpose:

Determining quantitative value of chloride in an unknown soluble salt by showing and explaining common methods in gravimetric analysis.

## Theory:

- Reaction between silver ion and chloride ion to form silver chloride as an equation:



- AgCl solubility in water and the  $K_{sp}$ :

Silver chloride is very insoluble, so we can assume that the above reaction goes to completion.

$$K_{sp} = 1.6 \times 10^{-10}$$

- AgCl is precipitated with an acidic solution to reduce likelihood of acid ions interfering with results. These ions will form co-precipitates with silver in the acid, due to these co-precipitates, the results would be higher.
- To prevent the precipitate from going through the filtered paper, it is heated to coagulate it into a different texture that's easier to work with. Heating the precipitate makes it coagulate, therefore, it is harder for the precipitate to go through the filtered paper.
- Dry silver chloride decomposes into silver and chlorine in the presence of light:



- If photodecomposition takes place in air, results will be lower, since chlorine will be released in the form of gas. If photochemical decomposition occurs in the presence of excess silver ions in aqueous solution, it will result in higher results.

# Unknown Salt Lab Report

**S Marginson**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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**English in Analytical Chemistry. Communicating about Methods & Techniques.** Книга для студента Надежда Зорина, Александра Соболева, 2022-10-19      *Lab Manual for Investigating Chemistry* David Collins (Ph. D.), Matthew

Johll, 2008-12-02 While many of the core labs from the first edition have been retained a renewed focus on the basics of chemistry and the scientific process create an even more detailed supplemental offering      **Modern Experimental**

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Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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Laboratory Manual for Principles of General Chemistry J. A. Beran,Mark Lassiter,2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section **Nuclear Science Abstracts** ,1975-08 **HIV Plus** ,2005-05 HIV Plus offers the latest stories on research economics and treatment The magazine raises awareness of HIV related cultural and policy developments in the United States and throughout the world **Chemistry in the Laboratory** James M. Postma,Julian L. Robert,J. Leland Hollenberg,2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

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