

Titration Practice Worksheet

Find the requested quantities in the following problems:

- 1) If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?

$$\begin{aligned} M_1 &= 0.1 \text{ M} \\ V_1 &= 54 \text{ mL} \\ M_2 &= ? \\ V_2 &= 125 \text{ mL} \end{aligned}$$

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (0.1 \text{ M})(54 \text{ mL}) &= (125 \text{ mL}) M_2 \\ M_2 &= \frac{(0.1 \text{ M})(54 \text{ mL})}{125 \text{ mL}} = 0.0432 \text{ M} = 4.32 \times 10^{-2} \text{ M} \end{aligned}$$

The concentration of HCl is $4.32 \times 10^{-2} \text{ M}$.

- 2) If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?

$$\begin{aligned} M_1 &= 0.05 \text{ M} \\ V_1 &= 25 \text{ mL} \\ M_2 &= ? \\ V_2 &= 345 \text{ mL} \end{aligned}$$

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (0.05 \text{ M})(25 \text{ mL}) &= (345 \text{ mL}) M_2 \\ M_2 &= \frac{(0.05 \text{ M})(25 \text{ mL})}{345 \text{ mL}} = 0.00362 \text{ M} = 3.62 \times 10^{-3} \text{ M} \end{aligned}$$

The concentration of NaOH is $3.62 \times 10^{-3} \text{ M}$.

- 3) If it takes 50 mL of 0.5 M KOH solution to completely neutralize 125 mL of sulfuric acid solution (H_2SO_4), what is the concentration of the H_2SO_4 solution?



$$n = MV$$

$$n = (0.5 \text{ M})(50 \text{ mL}) = 25 \text{ mol OH}^-$$

2 times $[\text{OH}^-]$ is needed

to neutralize $[\text{H}^+]$. $\rightarrow [\text{H}^+] = \frac{1}{2} [\text{OH}^-]$

$$M = \frac{12.5 \text{ mol}}{125 \text{ mL}} = \boxed{0.10 \text{ M}}$$

- 4) Can I titrate a solution of unknown concentration and still get a meaningful answer? Explain your answer in a few sentences.

Titration cannot be done without molarity of at least one of the substances. In order to solve $M_1 V_1 = M_2 V_2$, both molarities cannot be unknown.

- 5) Explain the difference between an endpoint and equivalence point in a titration.

Endpoint: titration can be stopped.

Equivalence point: when $[\text{H}^+]$ and $[\text{OH}^-]$ are neutralized.

The effectiveness of a titration is measured by the closeness of the endpoint to the equivalence point.

Solution Practice Answer Key

Thomas Bond, Chris Hughes



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