

1. What is a Colligative property?

Properties that depend on the concentration of solute particles but not on their identity. (For example: Freezing-point Depression, Boiling-point Elevation, and Osmotic Pressure)

2. What are freezing point depression, boiling point elevation, and osmotic pressure? List the formulas that are used to calculate each.

Freezing Point Depression (ΔT_f) – is the difference between the freezing points of the pure solvent and a solution of a nonelectrolyte in the solvent, and it is directly proportional to the molal concentration of the solution.

For an electrolyte solution, the number of ions (i) must be taken into account. ($\Delta T_f = i K_f m$)

Boiling-point Elevation (ΔT_b) – is the difference between the boiling points of the pure solvent and a nonelectrolyte solution of that solvent, and it is directly proportional to the molal concentration of the solution.

For an electrolyte solution, the number of ions (i) must be taken into account. ($\Delta T_b = i K_b m$)

Osmotic Pressure – is the external pressure that must be applied to stop osmosis. The number of ions (i) must also be taken into account in osmotic pressure calculations.

$$\pi = iMRT$$

3. What is the molal freezing-point constant (K_f) for water?

$$K_f = \frac{1.86\text{ }^{\circ}\text{C} \cdot \text{kg}}{\text{mole}}$$

4. What is the molal boiling point constant (K_b) for water?

$$K_b = \frac{0.52\text{ }^{\circ}\text{C}}{m}$$

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