

## Chapter 8 - Solution Stoichiometry - Multiple Choice Questions

- Solubility is a term:
  - that always means that more than 1g of solute will dissolve in 100 mL of water.
  - that can be used for all solutes because they all dissolve to some extent.
  - that generally means that at least 1g of solute will dissolve in 100 mL of water.
  - that is used in the same way by all chemists.
- Which statement is true?
  - When an atom loses an electron the cation formed is smaller and when an atom gains an electron the anion formed is larger.
  - When an atom gains or loses an electron the ion formed is smaller.
  - When an atom gains or loses an electron the ion formed is larger.
  - When an atom loses an electron the cation formed is larger and when an atom gains an electron the anion formed is smaller.
- Two important factors that affect the solubility of an ionic compound are:
  - ion charge and agitation.
  - ion size and agitation.
  - ion size and ion charge.
  - ion size and the particle size of the compound.
- Which statement is true?
  - Compounds with small ions tend to be more soluble than compound with large ions.
  - Compounds with large ions tend to be less soluble than compounds with small ions.
  - Compounds with small ionic charges tend to be less soluble than compounds with large ionic charges.
  - Compounds with large ions tend to be more soluble than compounds with small ions.
- Which anion when combined with  $Mg^{2+}$  will probably give the most soluble ionic compound?
  - $F^{-}$
  - $Cl^{-}$
  - $Br^{-}$
  - $I^{-}$
- Which cation when combined with oxide ion ( $O^{2-}$ ) is most likely to give a soluble compound?
  - $Ca^{2+}$
  - $K^{+}$
  - $Al^{3+}$
  - $Co^{2+}$
- According to the general solubility guidelines, which compound is not soluble?
  - $CaCO_3$
  - $LiBr$
  - $CaCl_2$
  - $Mg(NO_3)_2$
- According to the general solubility guidelines, which compound is soluble?
  - $BeSO_4$
  - $Na_2SO_4$
  - $Ag_2SO_4$
  - $CaSO_4$
- In a double displacement reaction:
  - two ionic compounds react by exchanging ions.
  - a precipitate may be formed.
  - a gas may be produced.
  - all of the above.
- Which combination of solutions will result in the formation of a precipitate?
  - $Ba(NO_3)_2(aq) + Na_2CO_3(aq)$
  - $KNO_3(aq) + Na_2CO_3(aq)$
  - $Ca(NO_3)_2(aq) + NaCl(aq)$
  - $NH_4NO_3(aq) + NaCl(aq)$
- A metal carbonate reacts with an acid to produce a gas. Which gas is formed?
  - $N_2(g)$
  - $CO_2(g)$
  - $CO(g)$
  - $H_2O(g)$
- Which of the following is one of the products of a reaction between an acid and a base?
  - Water
  - Hydronium ion
  - Hydroxide ion
  - Hydrogen gas
- Identify the spectator ions in the following total ionic equation:  
 $Na^{+}(aq) + OH^{-}(aq) + Ag^{+}(aq) + NO_3^{-}(aq) \rightarrow Na^{+}(aq) + AgOH(s) + NO_3^{-}(aq)$ 
  - $Na^{+}$  and  $OH^{-}$
  - $Ag^{+}$  and  $[NO_3]^{-}$
  - $OH^{-}$  and  $Ag^{+}$
  - $Na^{+}$  and  $[NO_3]^{-}$
- The following equation is an example of what type of equation?  
 $2Ag^{+}(aq) + S^{2-}(aq) \rightarrow Ag_2S(s)$ 
  - Molecular equation
  - Total ionic equation
  - Net ionic equation
  - Skeleton equation
- In writing the total ionic equation for the reaction between sodium phosphate and copper(II) chloride, which compound is not written as it's dissociated ions?
  - Sodium chloride
  - Copper(II) phosphate
  - Copper(II) chloride
  - Sodium phosphate
- The net ionic equation for the reaction that occurs from mixing a solution of barium chloride with a solution of potassium sulfate is:
  - $K^{+}(aq) + Cl^{-}(aq) \rightarrow KCl(aq)$
  - $BaCl_2(aq) \rightarrow Ba^{2+}(aq) + 2Cl^{-}(aq)$
  - $Be^{2+}(aq) + SO_4^{2-}(aq) \rightarrow BeSO_4(s)$
  - $BeSO_4(s) \rightarrow Be^{2+}(aq) + SO_4^{2-}(aq)$
- What is qualitative analysis?
  - It determines how much of a compound, element or ion is in an unknown sample.
  - It identifies compounds, elements or ions in an unknown solution.
  - It measures concentrations of ions.
  - None of the above.
- What is the dissociation equation for an aqueous solution of magnesium chloride?
  - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl_2(aq)$
  - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl^{-}(aq)$
  - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl_2^{-}(aq)$
  - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + 2Cl^{-}(aq)$
- What is the concentration of nitrate ions in a 0.15 mol/L solution of  $Fe(NO_3)_3$ ?
  - 0.45 mol/L
  - 0.30 mol/L
  - 0.15 mol/L
  - There is not enough information to answer the question.
- What is the concentration of a solution of aluminum sulfate ( $Al_2(SO_4)_3$ ) if the concentration of sulfate ions is 0.60 mol/L?
  - 1.2 mol/L
  - 0.60 mol/L
  - 0.40 mol/L
  - 0.20 mol/L
- Equal volumes of 0.20 mol/L  $NaCl$  and 0.30 mol/L  $KCl$  are mixed together. What is the concentration of chloride ions in the mixture?
  - 0.20 mol/L
  - 0.25 mol/L
  - 0.30 mol/L
  - You need to know the volumes of the original solutions to answer the question.
- What volume of a 0.500 mol/L solution of  $Na_2CO_3$  contains enough carbonate ions to completely precipitate 0.0550 mol of  $Be^{2+}$  ions?
  - 300 mL
  - 90.0 mL
  - 30.0 mL
  - 10.0 mL
- What volume of a 1.50 mol/L solution of  $Ba(OH)_2$  contains enough hydroxide ions to completely precipitate the iron(III) ions found in 60.0 mL of a 0.500 mol/L solution of  $FeCl_3$ ?
  - 180 mL
  - 60.0 mL
  - 20.0 mL
  - 12.5 mL
- How many moles of precipitate  $BeSO_4$  are formed when 50.0 mL of a 0.500 mol/L solution of  $BeCl_2$  is mixed with 35.0 mL of a 0.500 mol/L solution of  $Na_2SO_4$ ?
  - 0.0175 mol
  - 0.0250 mol
  - 0.0050 mol
  - 0.0800 mol
- How many moles of  $Mg(OH)_2$  will be precipitated when 75.0 mL of 0.400 mol/L  $MgCl_2$  is mixed with 75.0 mL of 0.400 mol/L  $NaOH$ ?
  - 0.0750 mol
  - 0.0600 mol
  - 0.0000 mol
  - 0.0050 mol
- In the above question, what mass of precipitate is formed?
  - 0.875 g
  - 1.75 g
  - 3.50 g
  - 4.37 g

# Solution Stoichiometry Multiple Choice Questions

**Barron's Educational Series, Neil D.  
Jespersen, Pamela Kerrigan**



### **Solution Stoichiometry Multiple Choice Questions:**

*Multiple Solution Methods for Teaching Science in the Classroom* Stephen DeMeo, 2008 For the first time in science education the subject of multiple solution methods is explored in book form While a multiple method teaching approach is utilized extensively in math education there are very few journal articles and no texts written on this topic in science Teaching multiple methods to science students in order to solve quantitative word problems is important for two reasons First it challenges the practice by teachers that one specific method should be used when solving problems Secondly it calls into question the belief that multiple methods would confuse students and retard their learning Using a case study approach and informed by research conducted by the author this book claims that providing students with a choice of methods as well as requiring additional methods as a way to validate results can be beneficial to student learning A close reading of the literature reveals that time spent on elucidating concepts rather than on algorithmic methodologies is a critical issue when trying to have students solve problems with understanding It is argued that conceptual understanding can be enhanced through the use of multiple methods in an environment where students can compare evaluate and verbally discuss competing methodologies through the facilitation of the instructor This book focuses on two very useful methods proportional reasoning PR and dimensional analysis DA These two methods are important because they can be used to solve a large number of problems in all of the four academic sciences biology chemistry physics and earth science This book concludes with a plan to integrate DA and PR into the academic science curriculum starting in late elementary school through to the introductory college level A challenge is presented to teachers as well as to textbook writers who rely on the single method paradigm to consider an alternative way to teach scientific problem solving

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