

## Stoichiometry: Multiple Choice

1. The limiting reactant of a chemical reaction is

- A) the reactant of a chemical process that is not consumed completely
- B) the product of a chemical process that is consumed completely
- C) the product of a chemical process that is not consumed completely
- D) the reactant of a chemical process that is consumed completely
- E) none of the above

ANS: D

1. Consider the following reaction:  $\text{Ca(OH)}_{2(aq)} + \text{Al}_2(\text{SO}_4)_{3(aq)} \rightarrow \text{CaSO}_{4(s)} + \text{Al(OH)}_{3(s)}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

- A) 1:1
- B) 2:1
- C) 2:3
- D) 1:2
- E) 3:1

ANS: E

1. Sodium phosphate,  $\text{Na}_3\text{PO}_4$ , can be prepared by combining sodium hydroxide and hydrogen phosphate as shown in the following balanced chemical equation:  $3\text{NaOH} + \text{H}_2\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{H}_2\text{O}$

If 2.5 mol of  $\text{H}_2\text{PO}_4$  are reacted, the number of moles of water that will be produced is

- A) 1.0
- B) 2.5
- C) 3.0
- D) 5.0
- E) 7.5

ANS: E

1. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

The number of moles of  $\text{KClO}_3$  required to produce 5.0 mol of oxygen is

- A) 1.0
- B) 2.0
- C) 3.0
- D) 3.3
- E) 5.0

ANS: D

1. Determine the number of grams of  $\text{CO}_2$  that can be produced from 48.0 g of the hydrocarbon  $\text{C}_2\text{H}_6$ , according to the following equation:



- A) 37.6 g
- B) 90.2 g
- C) 87.9 g
- D) 151 g
- E) 188.4 g

1. The limiting reactant in a chemical reaction is the \_\_\_\_\_ that is \_\_\_\_\_.

- A) Reactant, in excess
- B) Product, left over

# Solution Stoichiometry Multiple Choice Answer

**Albert A Gayle**



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