

## Chemical Equilibrium

**Ex 1.** For the equilibrium:  $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$

the equilibrium constant  $K_c$  is 50.5 at 448°C.

If  $[\text{H}_2]_0 = [\text{I}_2]_0 = 1 \text{ mol.L}^{-1}$ , what will the concentration of  $\text{H}_2$ ,  $\text{I}_2$  and  $\text{HI}$  be at equilibrium.

**Ex 2.** Consider the following reversible reaction:



One mole of  $\text{PCl}_5$  is placed in a flask of capacity 4 liter to be dissociated at constant temperature T. the number of moles of  $\text{Cl}_2$  formed at equilibrium is 0.25mole.

Calculate the equilibrium constant of the above system at temperature T.

**Ex 3.** At 100°C, the equilibrium constant for the reaction system,



is  $K_c = 2.19 \times 10^{-10}$

1. Determine whether the following mixture of components of the above system is at equilibrium:

$$[\text{COCl}_2] = 3.5 \times 10^{-3} \text{ mol.L}^{-1}$$

$$[\text{CO}] = 1.2 \times 10^{-5} \text{ mol.L}^{-1}$$

$$[\text{Cl}_2] = 3.3 \times 10^{-6} \text{ mol.L}^{-1}$$

2. If it is not at equilibrium, in what direction the system proceeds to reach equilibrium?

**Ex 4.** In a closed container of volume V and at constant temperature  $T = 300\text{K}$ , we introduce 2 mol of NO and 1 mol of  $\text{O}_2$  gases. the following equilibrium is established:



Calculate the degree of conversion ( $\alpha$ ) of NO if the amount of  $\text{NO}_2$  at equilibrium is three time that of  $\text{O}_2$ .

**Ex 5.** Ethanamine  $\text{C}_2\text{H}_5\text{NH}_2$  gas is put in an empty flask of volume  $V = 2\text{L}$ , maintained at constant temperature  $T = 500\text{K}$ .

The following equilibrium is established:



The initial amount of ethanamine is 0.2 mol and the total amount of the mixture at equilibrium is 0.3 mol.

1. Calculate the degree of dissociation  $\alpha$  of ethanamine.
2. Calculate the equilibrium constant  $K_c$  at 500K.

# Solutions Manual Chemical Equilibrium Chemistry

**Hans Kuhn, David H. Waldeck, Horst-  
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