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## Chapter 6 – Quantities in Chemical Reactions

- 6.1 (a) stoichiometry; (b) heat; (c) endothermic reaction; (d) specific heat; (e) actual yield
- 6.2 (a) limiting reactant; (b) exothermic reaction; (c) theoretical yield; (d) law of conservation of energy; (e) percent yield
- 6.3 (a) The balanced chemical equation for the metabolism of glucose is:  
$$\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$$
- (b) The coefficients of the balanced chemical equation give the proportions of the reactants and products in a reaction. From the balanced chemical equation, we see that the proportion of carbon dioxide to glucose is 6 molecules of  $\text{CO}_2$  to 1 molecule of  $\text{C}_6\text{H}_{12}\text{O}_6$ . When 12 molecules of  $\text{C}_6\text{H}_{12}\text{O}_6$  react, 72 ( $12 \times 6$ ) molecules of  $\text{CO}_2$  form:  
$$12 \text{ molecules } \text{C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ molecules } \text{CO}_2}{1 \text{ molecule } \text{C}_6\text{H}_{12}\text{O}_6} = 72 \text{ molecules } \text{CO}_2$$
- (c) From the balanced equation we see that glucose reacts with oxygen in a ratio of 1  $\text{C}_6\text{H}_{12}\text{O}_6$  molecule to 6  $\text{O}_2$  molecules. If 30 molecules of  $\text{O}_2$  react, then 5  $\text{C}_6\text{H}_{12}\text{O}_6$  molecules will react:  
$$30 \text{ molecules } \text{O}_2 \times \frac{1 \text{ molecule } \text{C}_6\text{H}_{12}\text{O}_6}{6 \text{ molecules } \text{C}_6\text{H}_{12}\text{O}_6} = 5 \text{ molecules } \text{C}_6\text{H}_{12}\text{O}_6$$
- 6.4 (a) The balanced equation for the reaction is:  
$$2\text{CH}_3\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 4\text{H}_2\text{O}$$
- (b) The coefficients of the balanced chemical equation give the proportions of the reactants and products in a reaction. From the balanced chemical equation, we see that 4 molecules of water react for every 2 molecules of methanol that react (a 2:1 ratio). If 10 molecules of methanol react, we expect 20 molecules of water to form:  
$$10 \text{ molecules } \text{CH}_3\text{OH} \times \frac{4 \text{ molecules } \text{H}_2\text{O}}{2 \text{ molecules } \text{CH}_3\text{OH}} = 20 \text{ molecules } \text{H}_2\text{O}$$

# Solutions Manual Chemical Reactions

**Lanny D. Schmidt**



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Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools

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Fundamentals of Chemical Reaction Engineering Mark E. Davis, Robert J. Davis, 2013-05-27 Appropriate for a one semester undergraduate or first year graduate course this text introduces the quantitative treatment of chemical reaction engineering It covers both homogeneous and heterogeneous reacting systems and examines chemical reaction engineering as well as chemical reactor engineering Each chapter contains numerous worked out problems and real world vignettes involving commercial applications a feature widely praised by reviewers and teachers 2003 edition

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