

Student Solution Manual

Student Solution Manual: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45055-0, Chapter 9: Solving Equilibrium Problems for Complex Systems

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

Chapter 9

- 9-2.** Why are simplifying assumptions in equilibrium problems restricted to relationships that are sums or differences?

Answer:

To simplify equilibrium calculations, we sometimes assume that the concentrations of one or more species are negligible and can be approximated as 0.00 M. In a sum or difference assuming a concentration is 0.00 M leads to an appropriate result. In contrast, if we were to simplify an equilibrium constant expression by assuming one or more concentrations are zero, we would be multiplying or dividing by 0.00, which would render the expression meaningless.

- 9-4.** Why do molar concentrations of some species appear as multiples in charge-balance equations?

Answer:

A charge-balance equation is derived by relating the concentration of cations and anions
no. mol/L positive charge = no. mol/L negative charge

For a doubly charged ion, such as Ba^{2+} , the concentration of charge for each mole is twice the molar concentration of the Ba^{2+} . That is,

$$\text{mol/L positive charge} = 2[\text{Ba}^{2+}]$$

Thus, the molar concentration of all multiply charged species is always multiplied by the charge in a charge-balance equation.

- 9-5.** Write the mass-balance expressions for a solution that is

Solutions:

- (a) 0.2 M in HF.

$$0.20 = [\text{HF}] + [\text{F}^-]$$

- (c) 0.10 M in H_3PO_4 .

$$0.10 = [\text{H}_3\text{PO}_4] + [\text{H}_2\text{PO}_4^-] + [\text{HPO}_4^{2-}] + [\text{PO}_4^{3-}]$$

- (e) 0.0500 M in HClO_2 and 0.100 M in NaClO_2 .

$$0.0500 + 0.100 = [\text{HClO}_2] + [\text{ClO}_2^-]$$

$$[\text{Na}^+] = C_{\text{NaClO}_2} = 0.100 \text{ M}$$

Solution Manual Skoog

**Douglas A. Skoog, Donald M. West, F.
James Holler, Stanley R. Crouch**



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the hosts and a range of microbes including bacteria and fungi Intercellular spaces vascular systems and even single cells can be inhabited by these endophytic microbes Of the bacterial endophytes only a small percentage is harmful to the plant most are neutral opportunistic or beneficial These plant based bacteria can have various important functions throughout the life cycle of the plant some promote plant growth and development others protect the plant from diseases This ability to be able to protect plants from diseases has catalyzed numerous laboratories to search for new bacteria that could be utilized instead of the traditional plant protective agents Because two or more interacting organisms are involved research and the eventual application of suitable bio controlling microbes are challenging and often require specific skills and equipment The purpose of this book is to provide a comprehensive review for those who are interested in the research and biotechnological applications of plant associated bacteria It also provides a compilation of current work conducted on plant bacteria interactions

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