

Student Solution Manual

Student Solution Manual: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45055-0, Chapter 2: Calculations Used in Analytical Chemistry

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

Chapter 2

2-1. Define

Answers:

(a) molar mass.

The molar mass is the mass in grams of one mole of a chemical species.

(c) millimolar mass.

The millimolar mass is the mass in grams of one millimole of a chemical species.

2-3. Give two examples of units derived from the fundamental base SI units.

Solution:

$$\text{The liter: } 1 \text{ L} = \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{\text{mL}} \times \left(\frac{\text{m}}{100 \text{ cm}} \right)^3 = 10^{-3} \text{ m}^3$$

$$\text{Molar concentration: } 1 \text{ M} = \frac{1 \text{ mol}}{\text{L}} \times \frac{\text{L}}{10^{-3} \text{ m}^3} = \frac{1 \text{ mol}}{10^{-3} \text{ m}^3}$$

2-4. Simplify the following quantities using a unit with an appropriate prefix

Solutions:

(a) $5.8 \times 10^8 \text{ Hz}$

$$5.8 \times 10^8 \text{ Hz} \times \frac{\text{MHz}}{10^6 \text{ Hz}} = 580 \text{ MHz}$$

(c) $9.31 \times 10^7 \mu\text{mol}$

$$9.31 \times 10^7 \mu\text{mol} \times \frac{\text{mol}}{10^6 \mu\text{mol}} = 93.1 \text{ mol}$$

(e) $3.96 \times 10^6 \text{ nm}$

$$3.96 \times 10^6 \text{ nm} \times \frac{\text{mm}}{10^6 \text{ nm}} = 3.96 \text{ mm}$$

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Henry Stark, John Woods, Jr.



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Skoog, 1998 *Solutions Manual for Principles of Instrumental Analysis, Third Edition* Douglas A. Skoog, 1985 **Student**

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Fundamentals of Analytical Chemistry Douglas A Skoog, Donald M West, F James Holler, Stanley R Crouch, 2021-03-18

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Kinsel, 2004 3 Using Spreadsheets in Analytical Chemistry 1 1 4 Calculations Used in Analytical Chemistry 2 12 5 Errors in Chemical Analyses 14 3 6 Random Errors in Chemical Analysis 17 8 7 Statistical Data Treatment and Evaluation 25 9 8 Sampling Standardization and Calibration 34 12 9 Aqueous Solutions and Chemical Equilibria 46 12 10 Electrolytes Effects on Chemical Equilibria 58 11 11 Solving Equilibrium Calculations for Complex Systems 69 9 12 Gravimetric Methods of Analysis 78 7 13 Titrimetric Methods Precipitation Titrimetry 85 12 14 Neutralization Titrations 97 20 15 Titration Curves for Complex Acid Base Systems 117 13 16 Applications of Neutralization Titrations 130 14 17 Complexation Formation and Precipitation Titrations 144 8 18 An Introduction to Electrochemistry 152 9 19 Applications of Standard Electrode Potentials 161 12 20 Applications of Oxidation Reduction Titrations 173 8 21 Potentiometry 181 10 22 Bulk Electrolysis Electrogravimetry and Coulometry 191 8 23 Voltammetry 199 4 24 Introduction to Spectrochemical Methods 203 5 25 Instruments for Optical Spectroscopy 208 3 26 Molecular Absorption Spectroscopy 211 9 27 Molecular Fluorescence Spectroscopy 220 3 28 Atomic Spectroscopy 223 5 29 Kinetic Methods of Analysis 228 6 30 An Introduction to Analytical Separations 234 7 31 Gas Chromatography 241 3 32 High Performance Liquid Chromatography 244 3 33 Miscellaneous Separation Methods 247 2 35 Preparing Samples for Analysis 249 1 36 Decomposing and Dissolving the Sample 250

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