

Chapter 1

1.1

Mass of water $= 10^5 \text{ kg}$, temperature raised by 20°C .

Heat needed $Q = 2 \times 10^7 \text{ cal} = 8.37 \times 10^7 \text{ J} = 23.2 \text{ kWh}$.

Work needed $= mgh = 14 \times 100 \times 20000 = 6.09 \times 10^7 \text{ ft-lb} = 22.9 \text{ kWh}$.

1.2

Work done along various paths are as follows

ab:

$$\int_a^b P dV = Nk_B T_1 \int_a^b \frac{dV}{V} = Nk_B T_1 \ln \frac{V_b}{V_a}$$

ac:

$$P_d(V_d - V_b) = Nk_B T_2 \left(1 - \frac{V_b}{V_d} \right)$$

dc:

$$Nk_B T_2 \int_d^c \frac{dV}{V} = Nk_B T_2 \ln \frac{V_c}{V_d}$$

No work is done along *bc* and *ca*. The total work done is the sum of the above. Heat absorbed equals total work done, since internal energy is unchanged in a closed cycle.

1.3

(a)

$$\alpha = \frac{1}{V} \frac{\partial V}{\partial T} = \frac{\partial V_0 T^{b-1}}{T_0^b V}$$

(b)

$$\Delta V = \frac{\partial V_0 T^{b-1}}{T_0^b} \Delta T$$

$$P = \frac{Nk_B T}{V} = \frac{Nk_B T_0^b}{V_0} T^{1-b}$$

$$\text{Work done} = P \Delta V = b N k_B \Delta T$$

Solutions Manual Chler Introduction To Statistical Mechanics

**Sylvain Ehrenfeld, Sebastian B. (Jt.
Author) Littauer**



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