

## Solutions, Acid/Base & Solubility Test Review

### Know the definitions for the following:

Solvent – liquid (usually water) that dissolves the solute to make the solution  
Solute – solid (usually) that is added to solvent and dissolves to make solution  
Solution – homogeneous mixture of solvent and solute  
Concentration – a measure of how much solute per volume of solvent  
Molarity – concentration; moles of solute/liters of solvent  
Arrhenius Acid and Base – acid donates  $H^+$  and base donates  $OH^-$   
Bronsted Lowry Acid and Base – acid donates  $H^+$  and a base accepts  $H^+$   
Neutralization – acid + base  $\rightarrow$  salt + water  
Titration – process used to determine concentration of an acid using neutralization

### Solutions/Solubility:

What factors will increase the rate of solubility? Increasing temperature; breaking into pieces; stirring

### Calculating Concentration (M = moles/liters)

What is the molarity of 50 grams of NaCl in a 100ml of water?

Convert grams  $\rightarrow$  moles: 50 grams NaCl / 58.5 grams (from periodic table) = .85 moles  
Divide moles by volume  $\rightarrow$  .85 moles / 0.1 L = 8.5 M

### Dilutions ( $M_1V_1 = M_2V_2$ )

If I have 500 mL of a 0.5M NaCl solution, what will the concentration be if the final volume of the solution is 600 mL?  $(500\text{ml}) \cdot (0.5\text{M}) = (600\text{ml}) ( ? \text{ M}) = 0.42 \text{ M}$

How many liters of 0.5M HCl stock solution would you need to prepare 2 liters of 0.25 M HCl solution?  $(2 \text{ L}) \cdot (0.25 \text{ M}) = (? \text{ L}) \cdot (0.5 \text{ M}) = 1 \text{ M}$

# Solutions Acid Base And Solubility Test Review Key

**JR Anderson**



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