

## Solutions, Acid/Base & Solubility Test Review

### Know the definitions for the following:

Solvent – liquid (usually water) that dissolves the solute to make the solution

Solute – solid (usually) that is added to solvent and dissolves to make solution

Solution – homogeneous mixture of solvent and solute

Concentration – a measure of how much solute per volume of solvent

Molarity – concentration; moles of solute/liters of solvent

Arrhenius Acid and Base – acid donates H<sup>+</sup> and base donates OH<sup>-</sup>

Bronsted Lowry Acid and Base – acid donates H<sup>+</sup> and a base accepts H<sup>+</sup>

Neutralization – acid + base → salt + water

Titration – process used to determine concentration of an acid using neutralization

### Solutions/Solubility:

What factors will increase the rate of solubility? Increasing temperature; breaking into pieces; stirring

### Calculating Concentration (M = moles/liters)

What is the molarity of 50 grams of NaCl in a 100ml of water?

Convert grams → moles: 50 grams NaCl / 58.5 grams (from periodic table) = .85 moles  
Divide moles by volume → .85 moles / 0.1 L = 8.5 M

### Dilutions (M<sub>1</sub>V<sub>1</sub> = M<sub>2</sub>V<sub>2</sub>)

If I have 500 mL of a 0.5M NaCl solution, what will the concentration be if the final volume of the solution is 600 mL? (500ml)\*(0.5M) = (600ml)( ? M) = 0.42 M

How many liters of 0.5M HCl stock solution would you need to prepare 2 liters of 0.25 M HCl solution? (2 L)\*(0.25 M) = (? L) \* (0.5 M) = 1 M

# Solutions Acid Base And Solubility Test Review Key

**JR Anderson**



## **Solutions Acid Base And Solubility Test Review Key:**

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