

Solutions Chemistry WebQuest

Directions:

- Go to <http://www.wwnorton.com/college/chemistry/gilbert2/contents/ch04/studyplan.asp>
- Under Chapter 4 ChemTours click the link that matches with the following headings.
- Read through the text on each page and watch the visuals shown. When finished, click the play button. If it asks you to click on anything, follow the instructions. Click on any *Concept Questions*, *Real-World Connections*, or *Science Connections* to get more information.

Molarity

1. Understanding the chemistry of liquid mixtures requires knowledge of amount or _____ of the compounds in the solution.
2. The concentration of solute is most usefully given in _____ ()
3. What is the equation for this?
4. How can you determine the number of moles of something?
5. Complete questions 1-7. Record your answers.
6. Calculating the molar concentration, or _____, of a solution is probably the _____ calculation done by _____ students and researchers.

Dilutions

7. Dilution occurs when additional _____ is added to the solution.
8. The _____ molecules are chemically unchanged during dilution.
9. When might you use a dilution? (read-world connections)
10. What is the equation used to solve dilution problems?
11. Both sides of the equation have exactly the same # of _____ of solute.
12. For solutions of the same compound in the same type of container, the _____ of the color is an indicator of solute _____.
13. The palest solution is the _____.
14. Answer questions 1& 2
15. Give 2 examples of when dilutions are used in science.

Migration of Ions

16. Some aqueous solutions will conduct _____ and some will not.
17. The flow of electricity is increased greatly when _____ ions travel to terminals & aid in _____ transfer.
18. _____ water is a poor conductor, but _____ or _____ water may contain ions that may cause electricity.
19. Why should you not go swimming outside during a thunderstorm?
20. In _____, ions can move freely and thereby conduct electricity.
21. What happens when a voltage is applied to Sodium Chloride Solution?
22. Why was the light bulb faint from the Acetic Acid solution?
23. Strong acid & bases, and ionic compounds are _____.
24. _____ & _____ are nonelectrolytes.
25. What types of ions would be found around the positive electrode?
26. What happens when 2 weak electrolytes are mixed?
27. Conductivity is another way to measure _____ in solution

Saturated Solutions

28. A saturated solution is one that contains the _____ concentration of _____.
29. A _____ is created when a substance is _____ in a _____.
30. When a solvent has reached its capacity and can dissolve no additional solute, it is said to be _____.
31. Give an example of a saturated solution.
32. What happens to compounds in an unsaturated solution?
33. In a saturated solution, excess molecules do not dissolve, but form a _____ on the bottom of the beaker.
34. What is happening to the molecule/ions in the beaker (section 3 of 4)?
35. What happens to solubility as temperature increases?

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1. Understanding the eBook Solutions Chemistry Webquest Key
 - The Rise of Digital Reading Solutions Chemistry Webquest Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Chemistry Webquest Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Chemistry Webquest Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Chemistry Webquest Key
 - Personalized Recommendations
 - Solutions Chemistry Webquest Key User Reviews and Ratings
 - Solutions Chemistry Webquest Key and Bestseller Lists
5. Accessing Solutions Chemistry Webquest Key Free and Paid eBooks
 - Solutions Chemistry Webquest Key Public Domain eBooks
 - Solutions Chemistry Webquest Key eBook Subscription Services
 - Solutions Chemistry Webquest Key Budget-Friendly Options
6. Navigating Solutions Chemistry Webquest Key eBook Formats

- ePub, PDF, MOBI, and More
- Solutions Chemistry Webquest Key Compatibility with Devices
- Solutions Chemistry Webquest Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solutions Chemistry Webquest Key
 - Highlighting and Note-Taking Solutions Chemistry Webquest Key
 - Interactive Elements Solutions Chemistry Webquest Key
- 8. Staying Engaged with Solutions Chemistry Webquest Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Chemistry Webquest Key
- 9. Balancing eBooks and Physical Books Solutions Chemistry Webquest Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Chemistry Webquest Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solutions Chemistry Webquest Key
 - Setting Reading Goals Solutions Chemistry Webquest Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solutions Chemistry Webquest Key
 - Fact-Checking eBook Content of Solutions Chemistry Webquest Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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