

PROBLEMS/SOLUTIONS

1. The mixing ratio of oxygen in the atmosphere is 20.95%. Calculate the concentration in mol L⁻¹ and in g m⁻³ at P° (101 325 Pa, 1.00 atm) and 25°C.

Solution

The mixing ratio for O₂ (g) is 20.95% (Table 2.1). Use the 'Ideal Gas' law (PV = nRT) to calculate the total number of moles of gas in 1.00 L for the given conditions. Constants are given in Appendices B.1 and C.1.

T = 25°C (298 K), P = 1.01325 x 10⁵ Pa, R = 8.315 J K⁻¹ mol⁻¹, and V = 1.00 L = 1.00 x 10⁻³ m³

The total number of moles of gas in 1.00 L is

$$\begin{aligned} n &= PV/RT \\ n &= 0.04089 \text{ mol} \end{aligned}$$

The number of moles of oxygen can be calculated from the mole fraction (which is the same as the mixing ratio).

$$\% \text{ O}_2 = 20.95, \text{ mole fraction} = 0.2095$$

The number of moles of O₂ in 1.00 L is: 0.2095 x 0.04089 = 0.00857 mol

The concentration of O₂ (g) is: 8.57 x 10⁻³ mol L⁻¹.

The concentration of O₂ (g) in units of g m⁻³ is determined as follows:

1 m³ = 1000 L, therefore, 1000 L contains 8.57 moles of O₂ (g). (M.M. of O₂ is 31.9988 g mol⁻¹)

$$31.9988 \text{ g mol}^{-1} \times 8.57 \text{ mol m}^{-3} = 274 \text{ g m}^{-3}$$

The O₂ (g) concentration can be expressed as either 8.57 x 10⁻³ mol L⁻¹ or 274 g m⁻³.

2. Calculate the atmospheric pressure at the stratopause. What are the concentrations (mol m⁻³) of dioxygen and dinitrogen at this altitude? How do these concentrations compare with the corresponding values at sea level?

Solution

Use Equation 2.3 to calculate the pressure at 50 km (stratopause).

$$P_h = P^\circ e^{-\bar{M}_a g}$$

R = 8.315 J K⁻¹ mol⁻¹, P° = 101 325 Pa, g = 9.81 m s⁻², $\bar{M}_a$ = 0.02896 kg mol⁻¹ (average molar mass of air), h = 50 km (50 000 m) and T = -2°C (271 K). Note: these last two parameters are the approximate altitude and temperature at the stratopause.

Solutions Manual For Environmental Chemistry

Ronald A. Hites, Jonathan D. Raff



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Hites, Jonathan D. Raff, 2013-08-26 From Reviews of the First Edition This splendid at times humorous and reasonably priced little book has much to commend it to undergraduate chemists and to other science students J G Farmer University of Edinburgh Complex environmental issues are presented in simple terms to help readers grasp the basics and solve relevant problems J Albaiges University of Barcelona The main strength of the book lies in its explanations of the calculation of quantitative relationships Each chapter includes 15-20 problems that are carefully chosen from a didactic standpoint for which the reader can find solutions at the end D Lenoir Institute for Ecological Chemistry What drew me to the first edition was the style the no nonsense down to earth explanations and the practical examples that litter the text The dry humor expressed in the footnotes is great and reminds me of other classic texts T Clough Lincoln University A practical approach to environmental chemistry Providing readers with the fundamentals of environmental chemistry and a toolbox for putting them into practice Elements of Environmental Chemistry Second Edition is a concise accessible and hands on volume designed for students and professionals working in the chemical and environmental sciences Tutorial in style this book fully incorporates real world problems and extensive end of chapter problem sets to immerse the reader in the field Chapters cover mass balance chemical kinetics carbon dioxide equilibria pesticide structures and much more Extensively revised updated and expanded this Second Edition includes new chapters on atmospheric chemistry climate change and polychlorinated biphenyls and dioxins and brominated flame retardants In addition new practice problems and a helpful tutorial on organic chemistry names and structures have been added to improve both the scope and accessibility of the book Principles and Applications of Aquatic Chemistry François M. M. Morel, Janet G. Hering, Laura Sigg, 2025-08-01 The definitive text for water chemistry professionals and students worldwide Principles and Applications of Aquatic Chemistry provides a solid foundation for understanding the chemistry of lakes oceans rivers estuaries and other natural waters Acclaimed for its user friendly pedagogy this classic textbook explains aquatic chemistry through the powerful application of the tableau system which provides a systematic way to organize complex chemical equilibrium problems Now in its second edition this title contains an entirely new introductory chapter and new coverage of ocean acidification advances in dissolution kinetics bioavailability of trace metals redox kinetics and updated thermodynamic data The use of computer programs to calculate chemical equilibrium in natural waters is illustrated Throughout this edition revised and streamlined material is supported by new real world examples and full color illustrations Accessible to those with diverse backgrounds in the sciences and engineering this essential textbook Covers the fundamentals of aquatic science including chemical thermodynamics acid base precipitation dissolution coordination reduction oxidation and adsorption reactions Explains the use of equilibrium calculations essential tools for understanding the chemical composition of aquatic systems and the fate of inorganic pollutants Provides quantitative treatments of the kinetics of chemical reactions in natural waters Features new and updated content that reflects advances in understanding the chemistry of natural waters Includes new end of chapter questions of various levels of

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