

Chemistry Study Guide

Periodic Table of the Elements

H 1 1.008																	He 2 4.003						
Li 3 6.941	Be 4 9.012																	B 5 10.811	C 6 12.011	N 7 14.007	O 8 15.999	F 9 18.998	Ne 10 20.180
Na 11 22.990	Mg 12 24.305																	Al 13 26.982	Si 14 28.086	P 15 30.974	S 16 32.065	Cl 17 35.453	Ar 18 39.948
K 19 39.098	Ca 20 40.078	Sc 21 44.956	Ti 22 47.883	V 23 50.942	Cr 24 51.996	Mn 25 54.938	Fe 26 55.845	Co 27 58.933	Ni 28 58.693	Cu 29 63.546	Zn 30 65.38	Ga 31 69.723	Ge 32 72.630	As 33 74.922	Se 34 78.96	Br 35 79.904	Kr 36 83.80						
Rb 37 85.468	Sr 38 87.62	Y 39 88.906	Zr 40 91.224	Nb 41 92.906	Mo 42 95.94	Tc 43 98.906	Ru 44 101.07	Rh 45 102.91	Pd 46 106.42	Ag 47 107.87	Cd 48 112.41	In 49 114.82	Sn 50 118.71	Sb 51 121.76	Te 52 127.6	I 53 126.91	Xe 54 131.29						
Cs 55 132.91	Ba 56 137.33																	Tl 81 204.38	Pb 82 207.2	Bi 83 208.98	Po 84 209	At 85 210	Rn 86 222
Fr 87 223	Ra 88 226																	Uut 113 288	Fl 114 289	Uup 115 288	Lv 116 293	Uus 117 286	Uuo 118 294
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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